

Pre-Lab Assignment – Lab 11 – Chemical Reactions

Answer: Chickadee's Notes

1. Match (draw a line) the following chemical reactions to their reaction types:



Combustion

Single-Displacement

Acid-Base

Double-Displacement

Decomposition

2. Briefly describe the actual process in each of the following reaction types:

Combustion: A combustion reaction is a reaction in which a hydrocarbon combines with oxygen (O_2) and releases energy, usually in the form of heat and light. Combustion reactions involve hydrocarbons.

Decomposition: These are the opposite of decomposition reactions. Thermally.

Thermally, in this sense, the heat induced decomposition. Electrolysis, electrolysis, is the process of decomposing through electricity. Photochemical, photochemical is the heat and light induced decomposition.

Single-Displacement: A single displacement reaction occurs when one element replaces another in a compound. This type of reaction has the general equation: $\text{A} + \text{BC} \rightarrow \text{B} + \text{AC}$. In this equation, A represents a metal element, element and BC represents the original compound.

Double-Displacement: In a double displacement reaction, two compounds react, often forming two different compounds. Like single displacement reactions, double displacement reactions always produce water and sometimes other ionic compounds. Double displacement reactions often produce precipitates when two solutions containing ions are mixed.

Acid-Base: Acid-base reactions occur between acids and bases. They are an example of neutralization reactions. They are also exothermic, with a heat of reaction, as they form water and salt.

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