

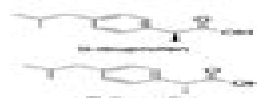
## Introduction to Chem 2443:

Organic chemistry = the branch of science that deals generally with compounds of carbon regardless of their origin

- Synthesis of organic compounds can be done in a lab

Inorganic = not present in living systems/organisms

- Molecular structure determines function!
  - Proteins – (1) structural and (2) active/enzymes
    - Fold because of hydrophobic parts that don't want to interact with water
  - 3D shape is very important
    - Solid line = coming towards
    - Dashed lines = going away



## Bonding and Hybridization:

- Atomic identity is based on the number of protons
  - Isotopes = different number of neutrons, but doesn't change the identity (protons)
- Bottom left = atomic number (protons=electrons)
- Upper left = atomic mass (protons + neutrons)

- Formal charges indicate when an atom has more or less valence electrons than what it should normally have



\*does not affect electronegativity

- Only valence electrons participate in chemical bonding
  - Sharing of V.E.s between atoms = covalent bonds
  - Covalent bonding is thermodynamically favored for stability
  - Hybrid orbitals are used for covalent bonding

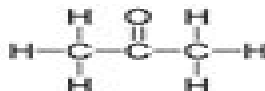
2 "things" →  
3 "things" →

Types of hybridization	Geometry
sp hybridization	Linear
sp <sup>2</sup> hybridization	Trigonal planar
sp <sup>3</sup> hybridization	Tetrahedral
dsp <sup>2</sup> hybridization	Square planar
dsp <sup>3</sup> hybridization	Trigonal bipyramidal
d <sup>2</sup> sp <sup>3</sup> hybridization	Octahedral
d <sup>2</sup> sp <sup>3</sup> hybridization	Pentagonal bipyramidal

"things" = lone pairs, single bond, double bond, triple bond

O always has 2 lone pairs, C has no lone pairs, N has one lone pair

O: sp<sup>2</sup>  
C: sp<sup>2</sup>



N: sp<sup>3</sup>



# Chem 2443 Organic Chemistry 9th Edition University Of Connecticut

**Forris Jewett Moore**



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