

Chapter 6 Worksheet



10. Compound F is known chemically to be the same as the acid, determine the empirical formula of compound F with the composition of 48.00% carbon, 4.37% hydrogen, and 47.63% oxygen.

$$\frac{48.00\% \text{ C}}{1} = \frac{\text{mass C}}{\text{mass F}} = 4.800 \text{ mol C}$$

$$\frac{4.37\% \text{ H}}{1} = \frac{\text{mass H}}{1.008 \text{ g}} = 4.34 \text{ mol H}$$

$$\frac{47.63\% \text{ O}}{1} = \frac{\text{mass O}}{\text{mass O}} = 2.975 \text{ mol O}$$

$\text{C}_{4.800} \text{H}_{4.34} \text{O}_{2.975}$



Emp. Form.
 $\text{C}_8 \text{H}_9 \text{O}_6$

11. A compound with molecular mass of 162 g/mol has the following composition: 65.05% carbon, 5.31% hydrogen, and 29.64% oxygen. Determine the molecular formula.

$$\frac{65.05\% \text{ C}}{1} = \frac{\text{mass C}}{12.01 \text{ g}} = 5.416 \text{ mol C}$$

$$\frac{5.31\% \text{ H}}{1} = \frac{\text{mass H}}{1.008 \text{ g}} = 5.267 \text{ mol H}$$

$$\frac{29.64\% \text{ O}}{1} = \frac{\text{mass O}}{16.00 \text{ g}} = 1.852 \text{ mol O}$$



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simplest
form

$\text{C}_{11} \text{H}_{11} \text{O}_4$
 is the
molecular
formula

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