

Thermodynamics
(Spontaneity and Entropy)

- T / F** A spontaneous process is one that proceeds on its on without any outside assistance.
- The total entropy of the universe _____ (increases/decreases) in any spontaneous process.
- When is entropy zero?
- Consider the vaporization of liquid water to steam at a pressure of 1 atm:
 - Is this process exothermic or endothermic?
 - In what temperature range is this process spontaneous?
 - What about nonspontaneous?
 - What temperature are the two phases in equilibrium?
- Of the following reactions determine if they will decrease/increase in entropy. Determine if ΔS will be positive, negative, or 0.
 - $\text{H}_2\text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(l)}$
 - $\text{Ag}^+_{(aq)} + \text{Cl}^-_{(aq)} \rightarrow \text{AgCl}_{(s)}$
 - $\text{N}_{2(g)} + \text{O}_{2(g)} \rightarrow 2\text{NO}_{(g)}$
 - $\text{CO}_{2(s)} \rightarrow \text{CO}_{2(g)}$
 - $\text{FeCl}_{3(s)} + \text{H}_{2(g)} \rightarrow \text{Fe}_{(s)} + 2\text{HCl}_{(g)}$
 - $\text{Ba(OH)}_{2(s)} \rightarrow \text{BaO}_{(s)} + \text{H}_2\text{O}_{(g)}$
- Calculate ΔS° values from the following reactions using the S° values from the table below:
 - $\text{N}_2\text{H}_{4(g)} + \text{H}_2_{(g)} \rightarrow 2\text{NH}_{3(g)}$
 - $2\text{Al}_{(s)} + 3\text{Cl}_{2(g)} \rightarrow 2\text{AlCl}_{3(g)}$
 - $\text{Mg(OH)}_{2(s)} + 2\text{HCl}_{(g)} \rightarrow \text{MgCl}_{2(s)} + 2\text{H}_2\text{O}_{(g)}$

	$\text{N}_2\text{H}_{4(g)}$	$\text{H}_{2(g)}$	$\text{NH}_{3(g)}$	$\text{Al}_{(s)}$	$\text{Cl}_{2(g)}$	$\text{AlCl}_{3(s)}$	$\text{Mg(OH)}_{2(s)}$	$\text{HCl}_{(g)}$	$\text{MgCl}_{2(s)}$	$\text{H}_2\text{O}_{(g)}$
J/molK	218.5	130.58	192.5	28.32	222.96	109.3	63.24	186.99	89.6	188.91

- The element of mercury, Hg, is a silvery liquid at room temperature. The normal freezing point of mercury is -38.9°C , and its molar enthalpy of fusion is $\Delta H_{\text{fus}} = 2.29 \text{ kJ/mol}$. What is the entropy change of the system when 50.0g of $\text{Hg}_{(l)}$ freezes at the normal freezing point?
- Choose the substance with the greater entropy in each case:
 - 1 mol $\text{H}_{2(g)}$ at STP or 1 mol of $\text{H}_{2(g)}$ at 100°C and 0.5 atm
 - 1 mol of $\text{H}_2\text{O}_{(l)}$ at 0°C or 1 mol of $\text{H}_2\text{O}_{(s)}$ at 25°C
 - 1 mol $\text{H}_{2(g)}$ at STP or 1 mol of $\text{SO}_{2(g)}$ at STP
 - 1 mol $\text{N}_2\text{O}_{(g)}$ at STP or 2 mol of $\text{NO}_{(g)}$ at STP

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