

Chemistry 1412: General Chemistry II - Lecture

Solutions Study Guide

Concepts

1. **Solutions** are homogeneous mixtures of two or more substances.
2. **solute**: substance present in smaller amount.
3. **solvent**: substance present in greater amount.
4. Liquid solutions are easiest to study because: (1) components are localized, (2) components mix well, and (3) they are easy to handle/transfer.
5. Solutions do not have to be liquid based.
6. **solubility**: amount of substance that will dissolve in a given amount of solvent.
7. Formation of a gas phase solution comprised of two noble gases is driven by **entropy**, not energy.
8. **entropy**: measure of energy of randomization or energy dispersal in a system.
9. Formation of liquid solutions are driven by **entropy** and **intermolecular forces**.
10. In solution, there are three types of particle interactions.
11. **“Like dissolves like” rule**: when solute and solvent have similar intermolecular forces, they tend to form a solution.
12. **miscible**: term used to describe two liquids that are soluble in all proportions.
13. The relationship between solvent-solute, solvent-solvent, and solute-solute interactions determines when a solution will form. If solvent-solute interactions are about equal to solvent-solvent and solute-solute interactions, entropy will decide whether a solution forms.
14. Energy changes are observed when solutions form. The **enthalpy of solution**, ΔH_{soln} , is the enthalpy associated with solution creation, and it can be estimated with a Hess's Law type calculation.
15. Formation of a solution can be broken up into three steps: (1) separation of solid into its constituent particles, ΔH_{solute} ; (2) separation of solvent particles to make room for solute, $\Delta H_{\text{solvent}}$; (3) mixing solute particles with solvent particles, ΔH_{mix} .
16. For ionic, water-based solutions, $\Delta H_{\text{solvent}}$ and ΔH_{mix} can be combined into $\Delta H_{\text{hydration}}$.

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Michael Seilmaier

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