

CHEM 101 Worksheet, Chapter 3: Mass Relationships in Chemical Reactions

Multiple choice questions: circle the correct answer(A-E)

Show your work questions (p. 2): show all work! Use the amu for each element from the periodic table!

$N_A = 6.02 \times 10^{23}$

1. Atom X has two isotopes: 25.0 wt% is ^{37}X and 75.0% is ^{35}X . What is the average molar mass of X?

- A. 36.0 amu B. 72.0 amu C. 36.5 amu D. 35.5 amu E. 34.5 amu

2. The molar mass (in amu) of $\text{ZnSO}_4 \cdot 6\text{H}_2\text{O}$ =

- A. 161.5 B. 179.5 C. 269.5 D. 53.5 E. 215.5

3. How many atoms of Na in 5.10 moles of Na_2CO_3 ?

- A. 3.07×10^{23} B. 6.14×10^{24} C. 6.14×10^{23} D. 1.84×10^{25} E. 10.2

4. What is the mass in grams of 1.00×10^{12} atoms of lead, Pb?

- A. $3.44 \times 10^{-10}\text{g}$ B. $2.07 \times 10^{114}\text{g}$ C. $1.25 \times 10^{38}\text{g}$ D. $3.44 \times 10^{-34}\text{g}$ E. $1.25 \times 10^{14}\text{g}$

5. 0.2 mol Na_2CO_3 = 4.24 g Na_2CO_3 = moles

- A. 530g, 449 mol B. 21.2g, 449 mol C. 530g, 0.051 mol D. 21.2g, 0.040 mol E. 16.6g, 0.051 mol

6. The wt% composition in tin(IV)oxide is %Sn %O (Sn = tin, element # 50)

- A. 21%Sn, 79%O B. 88%Sn, 12%O C. 12%Sn, 88%O D. 94%Sn, 6%O E. 79% Sn, 21%O

7. A compound has the following wt% composition: 2.1% H, 65.3% O, 32.6% S. What is its empirical formula?

- A. H_2SO_4 B. $\text{H}_4\text{S}_2\text{O}_8$ C. S_4HO_2 D. H_2SO_2 E. HO_4S

8. How many gram Na in 5.00g Na_2SO_4 ?

- A. 0.035 g B. 1.31 g C. 1.62 g D. 0.97 g E. 0.81 g

9. A compound containing only carbon and hydrogen has 85.7 wt% C and 14.3 wt% H. Its molar mass is 56 amu. What is its molecular formula?

- A. C_4H_8 B. C_2H_4 C. C_3H D. CH_2 E. CH

10. Balance the following reactions. No fractional coefficients, coefficients must be the lowest whole numbers.

a) K + H_2O \rightarrow KOH + H_2

b) H_2O_2 \rightarrow H_2O + O_2

c) C_2H_6 + O_2 \rightarrow CO_2 + H_2O

d) KOH + H_3PO_4 \rightarrow K_3PO_4 + H_2O

e) $\text{C}_2\text{H}_5\text{OH}$ + O_2 \rightarrow CO_2 + H_2O

11. How many g of CO_2 are formed from the combustion of 2.9 g of butane, C_4H_{10} , with excess oxygen?

The balanced reaction is: $2\text{C}_4\text{H}_{10} + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$

- A. 0.26g B. 6.6g C. 8.8g D. 10.5g E. 2.2g

12. How many g I_2 will react with 20.4 g Al to form AlI_3 (aluminum iodide)? As always, first balance the reaction! (Hint: this is not a limiting reagent problem)

- A. 144. g B. 288. G C. 192. G D. 72 g E. 576 g

13. For the balanced reaction: $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$. How many moles of NH_3 will be produced when 0.50 mol N_2 reacts with 1.0 mol H_2 ?

- A. 0.67 mol B. 1.0 mol C. 0.50 mol D. 2.0 mol E. 3.0 mol

14. Calculate the mass of CO_2 formed when 2.3 g of ethanol, $\text{C}_2\text{H}_5\text{OH}$ reacts with 3.2 g oxygen O_2 to form CO_2 and H_2O . (You balanced the reaction in 8.e)! This is a limiting reagent problem!

- A. 4.6g B. 2.2g C. 4.4g D. 6.6g E. 2.93g

Chem 101 Multiple Choice Questions

Wenbin Ji



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CliffsNotes AP Chemistry Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter's end Two full length practice tests with detailed answer explanations are included in the book

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