

Calorimetry Guided Notes

HS-PS3-4

CALORIMETRY GUIDED NOTES

Introduction to Calorimetry

Objectives

Students will understand the relationship between heat and temperature. They will be able to calculate the heat of reaction for a chemical reaction. They will be able to calculate the heat of formation for a compound. They will be able to calculate the heat of combustion for a fuel.

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Name _____ Date _____ Period _____

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What

Calorimetry is the study of heat. It is a branch of thermodynamics that deals with the measurement of heat. It is used to determine the heat of reaction for a chemical reaction. It is used to determine the heat of formation for a compound. It is used to determine the heat of combustion for a fuel.

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How to Use Calorimetry

Calorimetry is used to determine the heat of reaction for a chemical reaction. It is used to determine the heat of formation for a compound. It is used to determine the heat of combustion for a fuel.

1. Heat is the energy that is transferred from one system to another. It is measured in joules (J). The heat of reaction for a chemical reaction is the amount of heat that is released or absorbed during the reaction. The heat of formation for a compound is the amount of heat that is released or absorbed during the formation of the compound. The heat of combustion for a fuel is the amount of heat that is released during the combustion of the fuel.
2. Heat is the energy that is transferred from one system to another. It is measured in joules (J). The heat of reaction for a chemical reaction is the amount of heat that is released or absorbed during the reaction. The heat of formation for a compound is the amount of heat that is released or absorbed during the formation of the compound. The heat of combustion for a fuel is the amount of heat that is released during the combustion of the fuel.



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Calorimetry Guided Activity Student Guide

**Melinda S. Sothern, Stewart T.
Gordon, T. Kristian von Almen**



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