

A quantity of 85.0 mL of 0.900 M HCl is mixed with 85.0 mL of 0.900 M KOH in a constant-pressure calorimeter that has a heat capacity of 325 J/°C. If the initial temperatures of both solutions are the same at 18.24 °C, what is the final temperature of the mixed solution? The heat of neutralization is -56.2 kJ/mol. Assume the density and specific heat of the solutions are the same as those for water.



85.0 mL
0.900 M

85.0 mL
0.900 M

↓

↓

$$0.0765 \text{ mols} \times \frac{-56.2 \text{ kJ}}{1 \text{ mol}} = -4.2993 \text{ kJ}$$

← energy released to chemical react

$$C = 325 \frac{\text{J}}{^\circ\text{C}}$$

$$T_i = 18.24^\circ\text{C}$$

$$T_f = ?$$

$$S = 4.184 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}}$$

$$d = 1 \frac{\text{g}}{\text{mL}}$$

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