

Empirical Formulas and Molecular Formulas

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Activity 1

Remember, we already completed an empirical formula calculation using the mass percentages for oxygen and hydrogen for a compound. We have not accounted for all elements that are used in writing chemical formulas yet (carbon). This leads to the following information and we will handle elements

Table of the Chemical Data of Observed Compounds and Empirical Formulas

Compound	Empirical Formula	Observed Molar Mass	Calculated Molar Mass	Empirical Composition Factor	Empirical Formula Molar Mass
1	CH ₂ O	30.03	30.03		
2	C ₂ H ₄ O ₂	60.06	60.06		
3	C ₂ H ₄ O	44.05	44.05		
4	C ₂ H ₆ O	46.07	46.07		
5	C ₂ H ₆	30.07	30.07		

- Divide up the molar masses and calculated values; the integer represents the subscript for the particular element. (For the carbon, use the molar mass and calculated value)
- Identify the compound in Model 1 that has the same empirical formula.
- Identify the compound in Model 1 that has the same molar composition.
- Compare the compound in Model 1, together with compounds in the table, with different molecular formulas, that the same empirical formula has the same molar composition.
- Divide the molar masses in Model 1 (molar mass) by the empirical formula molar mass of Model 1 except the unit. Relationship of the molar mass of the given compound, the empirical formula of compound 1, and the given the compound's molecular formula.

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