

## Chapter 15: Acids and Bases



### Acids and Bases

#### Arrhenius Definitions:

- acids - compounds that produce an increase in  $[H^+]$  when dissolved in water
- bases - compounds that produce an increase in  $[OH^-]$  when dissolved in water

#### Lewis Definitions:

- acids - electron pair acceptors
- bases - electron pair donors

#### Bronsted-Lowry Definitions:

- acids -  $H^+$  donors
- bases -  $H^+$  acceptors

### Lewis Acids & Lewis Bases

- more broad way to define acids and bases
- Lewis acids – electron pair acceptors  
metal cations ( $M^{n+}$ ) and boron are common Lewis acids  
species that are electron deficient; electrophiles
- Lewis bases – electron pair donors  
species with O, N, halogen frequently have lone pairs of electrons to share  $\therefore$  Lewis bases  
species that are electron rich; nucleophiles
- product of a Lewis Acid + Lewis Base reaction is called a Lewis Acid-Base adduct

### Lewis Acids & Lewis Bases

- some examples:



- acidic oxides (oxides of nonmetals):



- basic oxides (oxides of metals):



# Chapter 15 Acids Bases Study Guide

**J Elliott**



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