

## Chapter 11 Alkynes

## Chapter Review

## General facts about alkynes

- Alkynes contain a carbon-carbon triple bond consisting of a strong  $\sigma$  bond and two weak  $\pi$  bonds. Each carbon is  $sp$  hybridized and linear (11.1).



- Alkynes are named using the suffix *-yne* (11.2).
- Alkynes have weak intermolecular forces, giving them low mp's and low bp's, and making them water insoluble (11.3).
- Because its weaker  $\pi$  bonds make an alkyne electron rich, alkynes undergo addition reactions with electrophiles (11.6).

## Addition reactions of alkynes

[1] Hydrohalogenation—Addition of  $\text{HX}$  ( $\text{X} = \text{Cl}, \text{Br}, \text{or I}$ ) (11.7)



- Markovnikov's rule is followed. H bonds to the *less* substituted C in order to form the more stable carbocation.

[2] Halogenation—Addition of  $\text{X}_2$  ( $\text{X} = \text{Cl or Br}$ ) (11.8)



- Anti addition of  $\text{X}_2$  occurs.

[3] Hydration—Addition of  $\text{H}_2\text{O}$  (11.9)



- Markovnikov's rule is followed. H bonds to the *less* substituted C in order to form the more stable carbocation.
- The unstable enol that is first formed rearranges to a carbonyl group.

Smith, J. G., & Smith, B. E. R. (2019). Chapter 11: Alkynes. In Student study guide/solutions manual to accompany Organic chemistry, sixth edition (pp 11/1-11/23). New York: McGraw-Hill Higher Education.

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