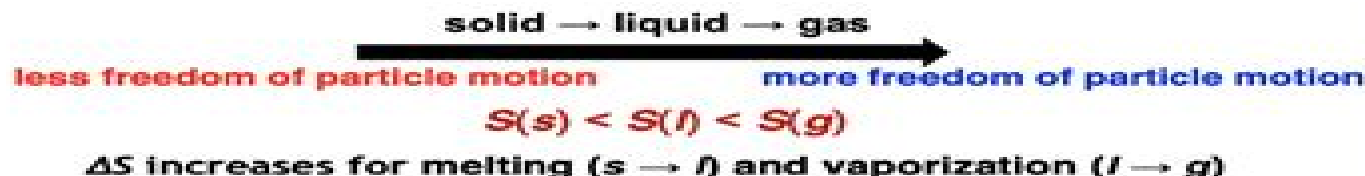


General Chemistry II Chapter 15 Study Guide

Entropy:

- Entropy (S): is related to freedom of motion of the particles in the system



Dissolving of soluble salts: solid + liquid \rightarrow ions in solution

localized energy of motion \rightarrow dispersed energy of motion

All spontaneous endothermic processes result in an increase
in *entropy* in the system.

Microstates and entropy:

- The entropy (S) of the system is also defined as the number of microstates (W) in a system
- $S = k \ln W$
 - k = Boltzmann constant (1.38×10^{-23} J/K)
 - W = # of microstates
- A system with fewer microstates has lower entropy; a system with more microstates has higher entropy
- Entropy, like enthalpy, is a state function and only depends on the initial and final states
 - $\Delta S = S_f - S_i$
 - S_f = final entropy
 - S_i = initial entropy

Spontaneous expansion of gas:

- When the stopcock is opened, the gas spontaneously expands to fill both flasks
- Increasing the volume increases the number of translational energy levels the particles can occupy; the number of microstates—and the entropy—increases

Factors affecting Entropy (S):

- Entropy of a substance is higher when:
 - Temperature of a substance is increased
 - # of substance particles is increased

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