



Name:

Date:

Student Exploration: Boyle's Law and Charles's Law

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: absolute zero, Boyle's law, Charles's law, Gay-Lussac's law, Kelvin scale, pressure

Prior Knowledge Question (Do this BEE OFE using the Gizmo.)

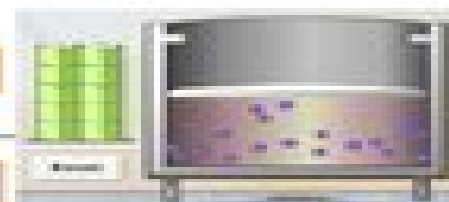
A small helium tank measures about two feet (60 cm) high. Yet it can fill over 50 balloons! How can such a small tank contain enough helium to fill so many balloons?

Gizmo Warm-up

The Boyle's Law and Charles's Law Gizmo shows a container of gas. Inside, small purple spheres represent gas molecules.

1. Observe the particles. Are they all moving at the same speed?

2. How do the particles interact with the walls and lid of the container?



These interactions contribute to the **pressure** on the walls of the container. Pressure is defined as force per unit area. The SI units of pressure are newtons per square meter (N/m^2), or pascals (Pa).

3. Slowly drag the temperature (**T**) slider back and forth. (Note: In this Gizmo, the **Kelvin scale** is used to measure temperature. On the Kelvin scale, 0 degrees is **absolute zero**, the coldest possible temperature. Absolute zero is equal to -273.15°C or -459.67°F .)

A. How does the change in temperature affect the speed of the molecules?

B. How does the change in temperature affect the volume of the container?

Chapter 14 Study Guide Chemistry Mixtures Solutions

Shasha Hu



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