

Chapter 10 Chemical Reactions

10.1 Law of Conservation of Matter

The Law of Conservation of Matter tells us that matter can never be created nor destroyed. This is very important in chemical reactions because it means that the mass of the reactants must equal the mass of the products. Or in other words, the number of atoms must equal the number of products atoms. Because of this law chemical reactions must be balanced.

What is a chemical equation? It is where we use formulas and symbols to describe a chemical reaction, or reactions, becoming products.



Synthesis: You must know which elements are solid, liquid, and gas??

sol = solid liq = liquid fl = liquid sol = no reaction A = heat
 (aq) = aqueous solution (ions or compounds dissolved in water)

Example: $\text{H}_2\text{SO}_4(\text{aq}) + 2\text{NaOH}(\text{aq}) \rightarrow \text{Na}_2\text{SO}_4(\text{aq}) + 2\text{H}_2\text{O}(\text{l}) + 2\text{OH}^-(\text{aq})$

Diatomic Molecules: You should know the seven elements that exist as **diatomic molecules**.

Hydrogen	Nitrogen	Fluorine	Oxygen	Carbon	Chlorine	Bromine
"H ₂ "	"N ₂ "	"F ₂ "	"O ₂ "	"C"	"Cl ₂ "	"Br ₂ "
"H ₂ "	"N ₂ "	"F ₂ "	"O ₂ "	"C"	"Cl ₂ "	"Br ₂ "

However these elements are not diatomic when they are in compounds. For example the correct formula for sodium chloride is NaCl not NaCl₂. These elements are only diatomic when they are alone as an element. Cl₂ has two Cl's because Cl is +2 charged so we need two +1 chlorines. The formula Cl is diatomic.

10.2 Balancing Chemical Equations

Due to the Law of Conservation, the number of atoms (and the mass) of each element must be equal on both sides of the equation. In other words, you must have the same number of each type of atom on both sides of the arrow. If there are four hydrogen atoms on the left side you must have four hydrogen atoms on the right side as well.

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M Carnoy



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