

Electrons in Atoms

Section 5.1 Light and Quantized Energy

pages 134–145

Practice Problems

pages 140–143

- Objects get their colors from reflecting only certain wavelengths when hit with white light. Light reflected from a green leaf is found to have a wavelength of 4.90×10^{-7} m. What is the frequency of the light?

$$c = \lambda \nu$$

$$3.00 \times 10^8 \text{ m/s} = (4.90 \times 10^{-7} \text{ m})\nu$$

$$\nu = \frac{3.00 \times 10^8 \text{ m/s}}{4.90 \times 10^{-7} \text{ m}} = 6.12 \times 10^{14} \text{ Hz}$$

- X rays can penetrate body tissues and are widely used to diagnose and treat disorders of internal body structures. What is the frequency of an X ray with a wavelength of 1.15×10^{-10} m?

$$c = \lambda \nu$$

$$3.00 \times 10^8 \text{ m/s} = (1.15 \times 10^{-10} \text{ m})\nu$$

$$\nu = \frac{3.00 \times 10^8 \text{ m/s}}{1.15 \times 10^{-10} \text{ m}} = 2.61 \times 10^{18} \text{ Hz}$$

- After careful analysis, an electromagnetic wave is found to have a frequency of 7.8×10^6 Hz. What is the speed of the wave?

$$3.00 \times 10^8 \text{ m/s}$$

- Challenge** While an FM radio station broadcasts at a frequency of 94.7 MHz, an AM station broadcasts at a frequency of 820 kHz. What are the wavelengths of the two broadcasts? Which of the drawings below correspond to the FM station? To the AM station?



FM station:

$$c = \lambda \nu$$

$$\lambda = c / \nu$$

$$\lambda = \frac{3.00 \times 10^8 \text{ m/s}}{9.47 \times 10^7 \text{ s}^{-1}} = 3.17 \text{ m}$$

FM: drawing b

AM station:

$$c = \lambda \nu$$

$$\lambda = c / \nu$$

$$\lambda = \frac{3.00 \times 10^8 \text{ m/s}}{8.20 \times 10^5 \text{ s}^{-1}} = 366 \text{ m}$$

The 820 kHz carrier wave has the greater wavelength.

AM: drawing a

- Calculate the energy possessed by a single photon of each of the following types of electromagnetic radiation.

a. $6.32 \times 10^{20} \text{ s}^{-1}$

$$E_{\text{photon}} = h\nu$$

$$= (6.626 \times 10^{-34} \text{ J} \cdot \text{s})(6.32 \times 10^{20} \text{ s}^{-1})$$

$$E_{\text{photon}} = 4.19 \times 10^{-13} \text{ J}$$

b. $9.50 \times 10^{13} \text{ Hz}$

$$E_{\text{photon}} = h\nu$$

$$= (6.626 \times 10^{-34} \text{ J} \cdot \text{s})(9.50 \times 10^{13} \text{ s}^{-1})$$

$$E_{\text{photon}} = 6.29 \times 10^{-20} \text{ J}$$

c. $1.05 \times 10^{16} \text{ s}^{-1}$

$$E_{\text{photon}} = h\nu$$

$$= (6.626 \times 10^{-34} \text{ J} \cdot \text{s})(1.05 \times 10^{16} \text{ s}^{-1})$$

$$E_{\text{photon}} = 6.96 \times 10^{-18} \text{ J}$$

Chapter 5 Electrons In Atoms Solution Manual

David R. Klein



Chapter 5 Electrons In Atoms Solution Manual:

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Brown, 2012-12-27 This is the student solutions manual to accompany Introduction to Organic Chemistry 5th Edition

Study Guide and Solutions Manual Neil E. Schore, 2002-08-02 *Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual* Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring

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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills,Amina El-Ashmawy,2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises Chemical Principles Study Guide/Solutions Manual John Krenos,Joseph Potenza,Carl Hoeger,2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications *Solutions Manual* Joseph Topich,2000-10

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