Electrons in Atoms

Section 5.1 Light and Quantized Energy

pages 134-145

Practice Problems

pages 140-143

 Objects get their colors from reflecting only certain wavelengths when hit with white light. Light reflected from a green leaf is found to have a wavelength of 4.90 × 10⁻⁷ m. What is the frequency of the light?

 $c = \lambda_F$ $3.00 \times 10^8 \text{ m/s} = (4.90 \times 10^{-7} \text{ m})_F$ $r = \frac{3.00 \times 10^8 \text{ pm/s}}{4.90 \times 10^{-7} \text{ pm}} = 6.12 \times 10^{14} \text{ Hz}$

 X rays can penetrate body tissues and are widely used to diagnose and treat disorders of internal body structures. What is the frequency of an X ray with a wavelength of 1.15 × 10⁻¹⁰ m?

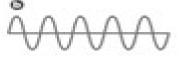
 $c = \lambda r$ $3.00 \times 10^8 \text{ ms/s} = (1.15 \times 10^{-10} \text{ ms})r$ $r = \frac{3.00 \times 10^8 \text{ ps/s}}{1.15 \times 10^{-10} \text{ ym}} = 2.61 \times 10^{18} \text{ Hz}$

 After careful analysis, an electromagnetic wave is found to have a frequency of 7.8 × 10⁶ Hz. What is the speed of the wave?

3.00 × 10⁸ m/s

4. Challenge While an FM radio station broadcasts at a frequency of 94.7 MHz, an AM station broadcasts at a frequency of 820 kHz. What are the wavelengths of the two broadcasts? Which of the drawings below correspond to the FM station? To the AM station?





FM station:

 $c = \lambda_F$ $\lambda = c/_F$

 $\lambda = \frac{3.00 \times 10^8 \text{ m/s}^2}{9.47 \times 10^3 \text{ s}^{-3}} = 3.17 \text{ m}$

FMI: drawing b

AM station:

 $c = \lambda \nu$ $\lambda = c/\nu$

 $\lambda = \frac{3.00 \times 10^8 \text{ m/s}}{8.20 \times 10^5 \text{ s}^{-7}} = 366 \text{ m}$

The 820 kHz carrier wave has the greater wavelength.

AM: drawing a

Calculate the energy possessed by a single photon of each of the following types of electromagnetic radiation.

a. $6.32 \times 10^{20} \, \mathrm{s}^{-1}$

 $E_{\text{photon}} = Prv$ = $(6.626 \times 10^{-34} \text{ J} \cdot \text{ph}(6.32 \times 10^{20} \text{ s}^{-4}))$

 $E_{\rm photon} = 4.19 \times 10^{-13} \, \text{J}$

b. $9.50 \times 10^{13} \, \mathrm{Hz}$

 $E_{\text{photon}} = hv$ = $(6.626 \times 10^{-34} \text{ J} \cdot \text{sh}(9.50 \times 10^{13} \text{ s}^{-34})$ $E_{\text{photon}} = 6.29 \times 10^{-20} \text{ J}$

c. $1.05 \times 10^{46} \text{ s}^{-1}$

 $E_{\text{photon}} = h_V$ = $(6.626 \times 10^{-34} \text{ J} \cdot \text{s})(1.05 \times 10^{16} \text{ s}^{-4})$

 $E_{\rm photon} = 6.96 \times 10^{-18} \, \text{J}$

×

Chapter 5 Electrons In Atoms Solution Manual

David R. Klein

Chapter 5 Electrons In Atoms Solution Manual:

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