



If the equilibrium constant is **very small** ($<10^{-3}$ for instance), then the reverse reaction is initially **very dominant** and the **reaction as written in the balanced equation** barely proceeds at all before equilibrium is reached.



$$K_c' = \frac{[\text{H}_2]^2 [\text{O}_2]}{[\text{H}_2\text{O}]^2} = 4.1 \times 10^{-48} \text{ at } 500 \text{ K}$$

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