

CHAPTER 15 STUDY GUIDE

Energy and Chemical Change

Section 15.1 Energy

In your textbook, read about the nature of energy.

In the space at the left, write *true* if the statement is true; if the statement is false, change the italicized word or phrase to make it true.

- | | |
|-------|---|
| _____ | 1. <i>Energy</i> is the ability to do work or produce heat. |
| _____ | 2. The law of conservation of energy states that energy <i>can be</i> created and destroyed. |
| _____ | 3. Chemical potential energy is energy stored in a substance because of its <i>composition</i> . |
| _____ | 4. <i>Heat</i> is a form of energy that flows from a warmer object to a cooler object. |
| _____ | 5. A calorie is the amount of energy required to raise the temperature of <i>one gram</i> of pure water by one degree Celsius. |
| _____ | 6. A <i>calorie</i> is the SI unit of heat and energy. |
| _____ | 7. The <i>specific</i> heat of a substance is the amount of heat required to raise the temperature of one gram of that substance by one degree Celsius. |
| _____ | 8. <i>Kinetic</i> energy is energy of motion. |
| _____ | 9. Chemicals participating in a chemical reaction contain only <i>potential</i> energy. |
| _____ | 10. One nutritional Calorie is equal to <i>100 calories</i> . |
| _____ | 11. One calorie equals <i>4.184 joules</i> . |
| _____ | 12. When a fuel is burned, some of its <i>chemical potential energy</i> is lost as heat. |
| _____ | 13. To convert kilojoules to joules, <i>divide</i> the number of kilojoules by 1000 joules/1 kilojoule. |

Answer the following question. Show all your work.

14. If the temperature of a 500.0-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J/(g·°C).

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Joacim Rocklöv



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Table of Contents Chapter 15 Study Guide Answer Key

1. Understanding the eBook Chapter 15 Study Guide Answer Key
 - The Rise of Digital Reading Chapter 15 Study Guide Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 15 Study Guide Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 15 Study Guide Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 15 Study Guide Answer Key
 - Personalized Recommendations
 - Chapter 15 Study Guide Answer Key User Reviews and Ratings
 - Chapter 15 Study Guide Answer Key and Bestseller Lists
5. Accessing Chapter 15 Study Guide Answer Key Free and Paid eBooks
 - Chapter 15 Study Guide Answer Key Public Domain eBooks
 - Chapter 15 Study Guide Answer Key eBook Subscription Services
 - Chapter 15 Study Guide Answer Key Budget-Friendly Options
6. Navigating Chapter 15 Study Guide Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 15 Study Guide Answer Key Compatibility with Devices
 - Chapter 15 Study Guide Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 15 Study Guide Answer Key
 - Highlighting and Note-Taking Chapter 15 Study Guide Answer Key
 - Interactive Elements Chapter 15 Study Guide Answer Key
8. Staying Engaged with Chapter 15 Study Guide Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chapter 15 Study Guide Answer Key
- 9. Balancing eBooks and Physical Books Chapter 15 Study Guide Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 15 Study Guide Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 15 Study Guide Answer Key
 - Setting Reading Goals Chapter 15 Study Guide Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 15 Study Guide Answer Key
 - Fact-Checking eBook Content of Chapter 15 Study Guide Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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