

## Chapter 24 Thermodynamics

## Exercises

## 24.1 Absolute Zero (pages 469)

1. Is the following sentence true or false? There is no limit to how cold an object can get. false

2. Define absolute zero.

Absolute zero is the temperature at which no more energy can be extracted from a substance, and no further lowering of its temperature is possible.

3. Circle the letter of each statement about a substance near absolute zero that is true.

- ☒ a. The thermal motion of its atoms approaches zero.  
☒ b. The kinetic energy of its atoms approaches zero.  
 c. A considerable amount of energy can still be removed from the substance.  
 d. Its temperature can still be significantly lowered.

4. Is the following sentence true or false? Negative Kelvin temperature values do not exist. true

5. Circle the letter that describes how the size of one Celsius degree and one Kelvin are related.

- ☒ a. They are equal.  
 b. Celsius degrees are smaller.  
 c. Celsius degrees are larger.  
 d. They have no consistent relationship.

Match each term or description to its Kelvin temperature.

| Term                                 | Definition |
|--------------------------------------|------------|
| <u>0</u> 6. absolute zero            | a. 373 K   |
| <u>-273</u> 7. melting point of ice  | b. 0 K     |
| <u>373</u> 8. boiling point of water | c. 273 K   |

## 24.2 First Law of Thermodynamics (pages 470–471)

9. Is the following sentence true or false? The flow of heat is not directly related to the flow of energy. false
10. The law of conservation of energy when applied to thermal systems is known as the first law of thermodynamics.
11. Circle the letter that best describes what happens when heat is added to a system.
- a. Much of it is destroyed immediately.  
☒ b. It transfers to an equal amount of some other form of energy.  
 c. Much of it is lost.  
 d. It is used to overcome friction.

# Chapter 24 Thermodynamics Exercises Answers

**Horia Metiu**



## Chapter 24 Thermodynamics Exercises Answers:

*Physical Chemistry: Thermodynamics* Horia Metiu, 2006-02-21 This is a new undergraduate textbook on physical chemistry by Horia Metiu published as four separate paperback volumes. These four volumes on physical chemistry combine a clear and thorough presentation of the theoretical and mathematical aspects of the subject with examples and applications drawn from current industrial and academic research. By using the computer to solve problems that include actual experimental data, the author is able to cover the subject matter at a practical level. The books closely integrate the theoretical chemistry being taught with industrial and laboratory practice. This approach enables the student to compare theoretical projections with experimental results, thereby providing a realistic grounding for future practicing chemists and engineers. Each volume of *Physical Chemistry* includes *Mathematica* and *Mathcad* Workbooks on CD-ROM. Metiu's four separate volumes *Thermodynamics*, *Statistical Mechanics*, *Kinetics*, and *Quantum Mechanics* offer built-in flexibility by allowing the subject to be covered in any order. These textbooks can be used to teach physical chemistry without a computer, but the experience is enriched substantially for those students who do learn how to read and write *Mathematica* or *Mathcad* programs. A TI-89 scientific calculator can be used to solve most of the exercises and problems.

**Perspectives on Statistical Thermodynamics** Yoshitsugu Oono, 2017-12-14 This original text develops a deep conceptual understanding of thermal physics, highlighting the important links between thermodynamics and statistical physics and examining how thermal physics fits within physics as a whole from an empirical perspective. The first part of the book is devoted to elementary mesoscopic topics such as Brownian motion, which leads to intuitive uses of large deviation theory, one of the pillars of modern probability theory. The book then introduces the key concepts behind statistical thermodynamics, and the final part describes more advanced and applied topics from thermal physics such as phase transitions and critical phenomena. This important subject is presented from a fresh perspective and in a highly pedagogical manner with numerous worked examples and relevant cultural side notes throughout, making it ideal as either a textbook for advanced thermal physics courses or for self-study by undergraduate and graduate students in physics and engineering.

**Solutions to Exercises** Roxy Wilson, 2002-05 Full solutions to all end-of-chapter exercises in the text are provided. With an instructor's permission, this manual may be made available to students.

**Student Solutions Manual to the Second Editions of Chemistry, Bailar ... [et al.] and Chemistry with Inorganic Qualitative Analysis, Moeller ... [et al.]** John Williams, 1984 *Solutions to Black Exercises* Roxy Wilson, 2005-05 Features detailed step-by-step solutions to the more than 1100 black-numbered end-of-character problems in *Chemistry: the central science*.

**Physical Chemistry** Peter Atkins, Julio de Paula, 2002 New edition of the overwhelmingly favorite text for the physical chemistry course.

**Handbook on Thermal Hydraulics in Water-Cooled Nuclear Reactors** Francesco D'Auria, Yassin A. Hassan, 2024-07-29 *Handbook on Thermal Hydraulics of Water Cooled Nuclear Reactors Volume 1: Foundations and Principles* includes all new chapters which delve deeper into the

topic adding context and practical examples to help readers apply learnings to their own setting Topics covered include experimental thermal hydraulics and instrumentation numerics scaling and containment in thermal hydraulics as well as a title dedicated to good practices in verification and validation This book will be a valuable reference for graduate and undergraduate students of nuclear or thermal engineering as well as researchers in nuclear thermal hydraulics and reactor technology engineers working in simulation and modeling of nuclear reactors and more In addition nuclear operators code developers and safety engineers will also benefit from the practical guidance provided Presents a comprehensive analysis on the connection between nuclear power and thermal hydraulics Includes end of chapter questions quizzes and exercises to confirm understanding and provides solutions in an appendix Covers applicable nuclear reactor safety considerations and design technology throughout Thermodynamics and the Free Energy of Chemical Substances Gilbert Newton

Lewis, Merle Randall, 1923 The scope of thermodynamics Definitions the concept of equilibrium Conventions and mathematical methods Solutions The first law of thermodynamics and the concept of energy The fugacity Application of the second law to solutions The perfect solution The laws of the dilute solution Systems involving variables other than pressure temperature and composition A useful function called the activity and its application to solutions Change of activity with the temperature and the calculation of activity from freezing points The standard change of free energy the equilibrium constant Solutions of electrolytes The activity of strong electrolytes The activity of electrolytes from freezing point data and tables of activity coefficients Activity coefficient in mixed electrolytes the principle of the ionic strength the activity of individual ions The galvanic cell Single potentials standard electrode potentials of the elements The third law of thermodynamics The entropy of monatomic gases and a table of atomic entropies Introduction to systematic free energy calculations the free energy of elementary hydrogen and metallic hydrides Oxygen and its compounds with hydrogen and with some metals Chlorine and its compounds Bromine and its compounds Iodine and its compounds Nitrogen compounds Carbon and some of its compounds Compounds of carbon and nitrogen Table of free energies and examples illustrating its use Conversion table for mol fractions mol ratios and molities Some useful numerical factors Coefficients employed in converting activity equilibrium constant and free energy from one temperature to another Publications by the authors pertaining to thermodynamics

*Solutions to Red Exercises - Chemistry* Roxy Wilson, 2005-05 Prepared by Roxy Wilson of University of Illinois Urbana Champaign Full solutions to all of the red numbered exercises in the text are provided Short answers to red exercises are found in the appendix of the text

**Solutions Manual to Accompany Elements of Physical Chemistry** C. A. Trapp, M. P. Cady, 2007 The Solutions manual to accompany Elements of Physical Chemistry 4e contains full worked solutions to all end of chapter exercises featured in the book

College Physics Textbook Equity Edition Volume 2 of 3: Chapters 13 - 24 An OER from Textbook Equity, 2016-02-11 This text is intended for one year introductory courses requiring algebra and some trigonometry but no calculus College Physics is organized such that topics are introduced conceptually

with a steady progression to precise definitions and analytical applications The analytical aspect problem solving is tied back to the conceptual before moving on to another topic Each introductory chapter for example opens with an engaging photograph relevant to the subject of the chapter and interesting applications that are easy for most students to visualize For manageability the original text is available in three volumes Original text published by Openstax College Rice University [www.textbookequity.org](http://www.textbookequity.org)

**Fundamentals of Physics, Part 3, Chapters 22 - 33, Enhanced Problems Version** David Halliday, Robert Resnick, 2002-04-16 The primary goal of this text is to provide students with a solid understanding of fundamental physics concepts and to help them apply this conceptual understanding to quantitative problem solving

Fundamentals of Physics, Chapters 22 - 45 David Halliday, Robert Resnick, Jearl Walker, 2001 The latest edition of Fundamentals of Physics has undergone a major redesign based on comments and suggestions from students and lecturers to make it more accessible to students and to provide them with an understanding of basic physics concepts

**Living by Chemistry (2018 Update)** Angelica M. Stacy, 2019-03-07 Designed to help all students to learn chemistry Living by Chemistry is a full year high school curriculum that incorporates science practices with a guided inquiry approach Students of all levels will gain a deep understanding of chemistry with this program With Living by Chemistry students learn chemistry in the same way that chemists work by asking questions collecting evidence and thinking like scientists Living by Chemistry is the product of a decade of research and development in high school classrooms focusing on optimizing student understanding of chemical principles Author Angelica Stacy assisted in the development of the NGSS standards and served on the AP Chemistry redesign committee She designed Living by Chemistry as an introduction for students who will take AP Chemistry or additional college classes The curriculum was developed with the belief that science is best learned through first hand experience and discussion with peers Guided inquiry allows students to actively participate in and become adept at scientific processes and communication These skills are vital to a students further success in science as well as beneficial to other pursuits Formal definitions and formulas are frequently introduced after students have explored scrutinized and developed a concept providing more effective instruction LBCs innovative curriculum offers much more than traditional programs To help engage students of all levels the curriculum provides a variety of learning experiences through activities discussions games demos lectures labs and individual work

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Fundamentals of Physics, Part 1, Chapters 1 - 12 David Halliday,Robert Resnick,2000-01-21 The primary goal of this text is to provide students with a solid understanding of fundamental physics concepts and to help them apply this conceptual understanding to quantitative problem solving **Thermodynamics and Chemistry, by F. H. MacDougall** Frank Henry Macdougall,1921 *Thermodynamics and Chemistry* Frank Henry MacDougall,1921 Fundamentals of Physics, A Student's Companion E-Book to Accompany Fundamentals of Physics, Enhanced Problems Version David Halliday,Robert Resnick,Jearl Walker,2003 A text for calculus based physics courses introducing fundamental physics concepts and featuring exercises designed to help students apply conceptual understanding to quantitative problem solving with chapter puzzlers checkpoints and reviews and summaries **Chemistry** Therald Moeller,2012-12-02 Chemistry with Inorganic Qualitative Analysis is a textbook that describes the application of the principles of equilibrium represented in qualitative analysis and the properties of ions arising from the reactions of the analysis This book reviews the chemistry of inorganic substances as the science of matter the units of measure used atoms atomic structure thermochemistry nuclear chemistry molecules and

ions in action This text also describes the chemical bonds the representative elements the changes of state water and the hydrosphere which also covers water pollution and water purification Water purification occurs in nature through the usual water cycle and by the action of microorganisms The air flushes dissolved gases and volatile pollutants when water seeps through the soil it filters solids as they settle in the bottom of placid lakes Microorganisms break down large organic molecules containing mostly carbon hydrogen nitrogen oxygen sulfur or phosphorus into harmless molecules and ions This text notes that natural purification occurs if the level of contaminants is not so excessive This textbook is suitable for both chemistry teachers and students

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