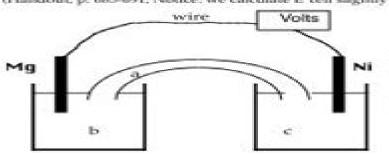
Chapter 23: Electrochemistry Voltaic Cells, Cell Potentials (p. 677-681) Page 4 (Handout, p. 685-691, Notice: we calculate E'cell slightly differently than the book!)



- Consider this voltaic cell made with Mg and Ni:
- a. part "a" is called the
- b. What solution is in container "b"?
 - Mg(NO₃)₂ or Ni(NO₃)₂
- c. What solution is in container "c"?
 - Mg(NO₃)₂ or Ni(NO₃)₂
- d. Does the oxidation or reduction occur at the Mg electrode? Is Mg the red wire or black wire?

Write the equation for this half reaction _____

e. Does the oxidation or reduction occur at the Ni electrode? Is Ni the red wire or black wire?

Write the equation for this half reaction _____

- f. Write the overall reaction for the cell _____
- g. Which way do electrons flow in the wire? _____
- h. Calculate Eⁿoett for this cell.
- a. For the Fe⁺Fe⁺ and Cu⁺Cu⁺ half cells: Which is oxidized?____ Which is reduced? ____
 Write the half reactions for the oxidation and the reduction and calculate E⁺ort.
- b. For the Al | Al of and Ag | Ag half cells: Which is exidized?_____ Which is reduced? _____ Which is reduced? _____
 Write the half reactions for the exidation and the reduction and calculate E of the control of the exidation.
- A voltaic cell is made with Paladium (Pd | Pd = 2) = red wire and I2 | I = black wire. The measured E online is + 0.41 Volts. I2 | I is found on your E od chart. Pd | Pd = 2 is not.
 - a. Is Pd | Pd+2 the oxidation or the reduction? Is I₂ | I⁻ the oxidation or the reduction?
 - b. Calculate E^o_{red} for the Pd⁻² + 2 e⁻—> Pd reaction using this data.
 - c. Determine E°ox for Pd ---> Pd+2 + 2 e-

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Petrogav International Oil & Gas Training Center

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