

THERMOCHEMISTRY

Q1: Define the following terms: Thermochemistry, joule, calorie, exothermic reactions, endothermic reactions, system surrounding and boundaries of the system.

Ans Thermochemistry: (ETEA: 2005)

"The branch of chemistry which deals with the study of amount of heat evolved or absorbed during a chemical reaction is called thermochemistry."

Joule:

The SI unit of energy is joule. "The energy expended when one Newton force, displaces a body through one meter is called one joule energy."

Calori:

One calorie is the amount of heat, energy needed to raise the temperature of one gram of water from 14.5°C - 15.5°C.

$$1 \text{ Cal} = 4.18 \text{ joule}$$

$$1 \text{ K-Cal} = 1000 \text{ Cal}$$

Thermochemical Reactions:

"Chemical reactions during which heat is either evolved or absorbed are called thermochemical reactions." Thermochemical reactions are of two types:

1. Exothermic reactions
2. Endothermic reactions

1. Exothermic Reaction:

"Exo = outside, therm = heat"

"Chemical reactions during which heat is given out are called exothermic reactions." The released energy is represented by negative sign.

Examples:



$$\Delta H = -296.6 \text{ KJ}$$



$$\Delta H = -92.3 \text{ KJ}$$

Chapter 11 Thermochemistry Guided Reading Answers

**Charles William Keenan, Donald C.
Kleinfelter, Jesse Hermon Wood**



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 - Chapter 3: Chapter 11 Thermochemistry Guided Reading Answers in Everyday Life
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