

## TEACHER GUIDE AND ANSWERS

### Study Guide - Chapter 10 – The Mole

#### Section 10.1 Measuring Matter

1. pair
2. 5
3. dozen
4. gross
5. 200
6. ream
7. 6,000,000,000
8. 0.5 mol
9.  $6.02 \times 10^{23}$
10. four moles
11.  $\frac{6.02 \times 10^{23} \text{ Cu atoms}}{1 \text{ mol Cu}}$
12.  $\frac{1 \text{ mol CH}_4}{6.02 \times 10^{23} \text{ molecules CH}_4}$
13.  $\frac{1 \text{ mol Xe}}{6.02 \times 10^{23} \text{ molecules Xe}}$
14.  $\frac{6.02 \times 10^{23} \text{ molecules F}_2}{1 \text{ mol F}_2}$

#### Section 10.2 Mass and the Mole

1. false
2. true
3. false
4. true
5. true
6. true
7. false
8. true
9. b
10. d, a
11. a
12. b, c
13. d
14. c

#### Section 10.3 Moles of Compounds

1. carbon, 1 atom; hydrogen, 4 atoms
2. carbon, 1 atom; hydrogen, 1 atom; chlorine, 3 atoms
3. 1 mol C; 4 mol H
4. 1 mol C; 1 mol H; 3 mol Cl
5.  $6.02 \times 10^{23}$
6.  $1.81 \times 10^{24}$
7. 16.05 g/mol
8.  $1.20 \times 10^{24}$

#### Section 10.4 Empirical and Molecular Formulas

1. The percent composition of a compound is the percent by mass of each of the elements in a compound.
2. Divide the mass of each element in the sample by the mass of the sample. Then multiply each quotient by 100.
3. c
4. c
5. b
6. b
7. c
8. c
9. d
10.  $7.89 \text{ g K} \times 1 \text{ mol K} / 39.10 \text{ g K} = 0.202 \text{ mol K}$   
 $2.42 \text{ g C} \times 1 \text{ mol C} / 12.01 \text{ g C} = 0.201 \text{ mol C}$   
 $9.69 \text{ g O} \times 1 \text{ mol O} / 16.00 \text{ g O} = 0.606 \text{ mol O}$   
 $0.202 \text{ mol K} / 0.201 \text{ mol C} = 1.00 \text{ mol K} / 1.00 \text{ mol C} = 1 \text{ mol K} / 1 \text{ mol C}$   
 $0.201 \text{ mol C} / 0.201 \text{ mol C} = 1.00 \text{ mol C} / 1.00 \text{ mol C} = 1 \text{ mol C} / 1 \text{ mol C}$   
 $0.606 \text{ mol O} / 0.201 \text{ mol C} = 3.01 \text{ mol O} / 1.00 \text{ mol C} = 3 \text{ mol O} / 1 \text{ mol C}$   
1 mol K : 1 mol C : 3 mol O  
empirical formula:  $\text{KCO}_3$   
(39.10 g + 12.01 g + 48.00 g) = 99.11 g

# Chapter 10 The Mole Study Guide Answers

**David G. Myers, Richard O. Straub**



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