Soft X-Rays and Extreme Ultraviolet Radiation: Principles and Applications

Chapter 1 Homework Problems

1.1 (a) Confirm that for a photon energy $\hbar \omega$, with corresponding wavelength λ in vacuum, the two are related by

$$hos \cdot \lambda = 1239.842 \text{ nm} \cdot \text{eV}$$

Values for physical constants are given in Appendix A, page 419. (b) What wavelength in nm corresponds to photons of energy 543.1 eV? (c) What wavelengths in vacuum are associated with the K and L-absorption edges of neon, chlorine, cobalt, and ruthenium?

1.2 (a) Show that for a wavelength λ in vacuum, the number of photons N_p per joule of energy E is given by

$$N_g = 5.034 \times 10^{13} \text{ photons} \left(\frac{\lambda}{\text{nm}}\right) \left(\frac{E}{\text{joules}}\right)$$

(b) What number of photons is required for a total energy of 1 μJ at a vacuum wavelength of 530 nm (green light); 13.4 nm (EUV); 2.28 nm (soft x-ray); and 0.0620 nm (hard x-ray)?

1.3 Show that the photon flux F per watt of radiated power P can be written as

$$F = (5.034 \times 10^{11} \text{ ph/s}) \left(\frac{\lambda}{\text{nm}}\right) \left(\frac{P}{\text{watt}}\right)$$

where λ is the wavelength in vacuum.

1.4 If Cu K_{0.1} x-rays irradiate an aluminum foil, (a) with what energy in eV will K-shell photoelectrons be ejected from the Al atoms? (b) What characteristic lines of Al might subsequently be observed? (c) Which Auger electrons might be observed? Estimate the energy of a KL₂L₃ Auger electron. Hint: The binding energy of the L₃ electron with a vacancy in the K-shell is not exactly the same as the binding energy of the L₃ electron in the natural state. For further information on measurements of Auger electrons see: http://srdata.nist.gov/xps/main_search_menu.htm. (d) How would it be possible to differentiate photoelectrons from Auger electrons, in a more general case, when investigating materials with photons?

1.5 (a) What is the K-absorption edge of carbon in eV? (b) What is the corresponding wavelength in vacuum? (c) What is the photon energy of K_n emissions from a carbon atom? (d) Repeat this exercise for oxygen and iron.

1.6 (a) Compute and graph the transmission curves (linear transmission, log photon energy; 0-100% and 100 eV to 10 keV, respectively) for (a) 15 μm thick aluminum foil and (b) 75 μm thick Saran wrap (two common wraps for sandwiches, also common as x-ray filters). For Al use ρ = 2.70 g/cm³; for Saran wrap (CH₂=CCl₂) use ρ = 1.60 g/cm³ (total).

1.7 (a) Calculate the atomic densities of aluminum and silicon based on mass density and atomic weight values given in the Periodic Table of the Elements (inside back cover of this book). (b) Estimate the corresponding interatomic distances based on a geometric model. Compare these to the "nearest neighbor" values given in the Periodic Table. Refer to p.440 in the text by R. Leighton for further discussion (reference 27 of this chapter).

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United States. National Bureau of Standards

E Study Guide For Soft X Rays And Extreme Ultraviolet Radiation:

Soft X-Rays and Extreme Ultraviolet Radiation David Attwood,2000 This comprehensive book describes the fundamental properties of soft X rays and extreme ultraviolet EUV radiation and discusses their applications in a wide variety of fields including EUV lithography for semiconductor chip manufacture and soft X ray biomicroscopy The author provides a wealth of useful reference material such as electron binding energies characteristic emission lines and photo absorption cross sections The book will be of great interest to graduate students and researchers in engineering physics chemistry and the

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