

**POGIL: Electron Configuration and Orbitals****Model 1: Orbitals**

The quantum mechanical model determines the allowed energies an electron can have and how likely it is to find the electron in various locations around the nucleus. Electrons are located in orbitals (different than Bohr!) in the electron cloud. An **atomic orbital** is a region of space in which there is a high probability of finding an electron. There are four orbitals, s, p, d, and f which are found in different energy levels called **principle quantum numbers**. Below is a table summarizing this idea. Be aware that we go all the way up to principle energy level of 7, however, it does become a bit more complex. We use the energy level numbers and orbitals in a "code" called **electron configuration** which helps us express the location of an electron with high probability.

Summary of Principal Energy Levels, Sublevels, and Orbitals		
Principal energy level	Number of sublevels	Type of sublevel
$n = 1$	1	1s (1 orbital)
$n = 2$	2	2s (1 orbital), 2p (3 orbitals)
$n = 3$	3	3s (1 orbital), 3p (3 orbitals), 3d (5 orbitals)
$n = 4$	4	4s (1 orbital), 4p (3 orbitals), 4d (5 orbitals), 4f (7 orbitals)

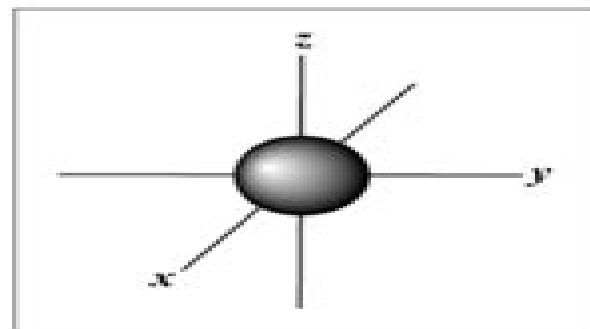
**Key Content Questions:**

1. What is an atomic orbital? What are the four different orbitals?
2. What do you notice about the sublevels as you increase in energy levels?
3. Are all orbitals found in each sublevel?

**s Orbital**

The s orbital is a spherical shape and can only have 2 electrons. It is the first orbital in an energy level to be filled when writing electron configurations. The s orbital only has one orientation, seen in the picture to the right. This orbital is found in all energy levels.

(s=sphere...memory clue!)



# Dynamic Equilibrium Which Way Do We Go Pogil Answers

**Arturo Cuomo**



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