

CH 15 Equilibrium Study Guide

Dynamic Equilibrium

Dynamic Equilibrium – forward & reverse reaction rates are equal

- **ASSUME** all rxns are **ELEMENTAL** (1 step!)
- Reversible reaction - something constantly produced & consumed
 - Indicated by \rightleftharpoons

No net change in reactant or product concentrations

Equilibrium Constant, K_c

For a general equation:



$$K_c = \frac{k_f}{k_r} = \frac{[D]^d [E]^e}{[A]^a [B]^b}$$

- A. The equilibrium-constant depends only on the stoichiometry of the reaction, not the mechanism.
- B. **DOES NOT** depend on initial concentrations because *only* the **RATIO** of concentrations at equilibrium is the same
- C. **K** has **NO** units!
- D. **ONLY** consider gas & aqueous phases!!
-OMIT solids & liquids

Significance of K

Value of K indicates if more products or reactants are present *at equilibrium*

- If K is large (>1): Products favored - Equilibrium lies to the right
- If K is small (<1): Reactants favored - Equilibrium lies to the left

Equilibrium Constant, K_p

For a general equation:

Only if ALL phases are gases!



Relating K_p and K_c

$$K_p = K_c (RT)^{\Delta n} ; \Delta n = [\sum (\text{moles GAS in product}) - \sum (\text{moles GAS in reactant})]$$

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