

- 5.) 350.0 liters of argon gas is held at -165°C and 6.64 atm. What temperature will this gas have if it exerts 5.50 atm in a 275.0 liters tank? (3 pts.)

$$\frac{350.0\text{ L} \cdot 6.64\text{ atm}}{108^{\circ}\text{K}} = \frac{275.0\text{ L} \cdot 5.50\text{ atm}}{X}$$

$$103350 = 2324X$$

$$X = 44.47^{\circ}\text{K}$$

- 6.) 22.0ml of a gas is held at STP. What volume will it occupy at -30.0°C and 570 torr? (3 pts.)

$$\frac{22.0\text{ mL} \cdot 760\text{ torr}}{273^{\circ}\text{K}} = \frac{X \cdot 570\text{ torr}}{243^{\circ}\text{K}}$$

$$4662960 = 155610X$$

$$X = 30.0\text{ mL}$$

- 7.) 3.266 moles of phosphine gas occupies 25.0 liters at 2440 torr and 300.K. If 1.141 moles of the phosphine gas are lost, what temperature will the gas have to have in order to keep 20.0 liters and a pressure of 3.00×10^3 torr? (4 pts.)

$$\frac{25.0\text{ L} \cdot 2440}{3.266 \cdot 300\text{ K}} = \frac{20.0\text{ L} \cdot 3.00 \times 10^3}{2.125 \cdot X}$$

$$129625X = 58788000$$

$$X = 454.5\text{ K}$$

$$181.0^{\circ}\text{C}$$

- 8.) A sample of a gas has a density of 2.17 g/L at 1.42 atm and 20.0°C . What will the density of this gas be at -30.0°C and 2.17 atm? (4 pts.)

$$\frac{2.17\text{ g/L} \cdot 1.42\text{ atm}}{1\text{ mole} \cdot 273} = \frac{X \cdot 2.17}{1\text{ mole} \cdot 243}$$

$$5874361.74 = 258133X$$

$$X = 22.75\text{ g/L}$$

- 9.) 400.0 ml of a gas exists at a certain temperature and pressure with moles held constant. What volume (in ml) will this gas occupy if its absolute temperature is reduced by a third and the pressure is doubled? (5 pts.)

$$400.0\text{ mL} \left(\frac{T}{\frac{1}{3}} \right) \left(\frac{P}{2} \right) = 66.67\text{ mL}$$

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