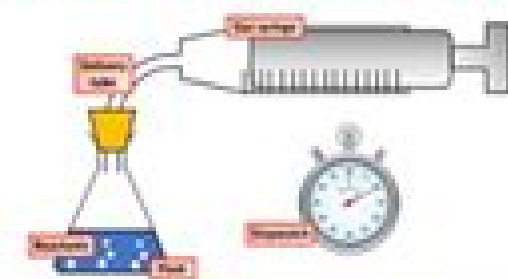


## Calculating Rate

$$\text{Rate} = \frac{\text{amount of product made}}{\text{Time taken}}$$

## Investigating Rate of Reaction

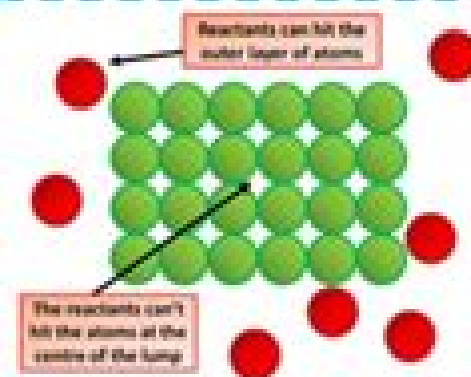
1. Support a gas syringe with a stand, boss and clamp.
2. Using a measuring cylinder, add 10 cm<sup>3</sup> of dilute hydrochloric acid to a conical flask.
3. Add 0.4 g of calcium carbonate to the flask, immediately connect the gas syringe and start a stop clock.
4. Every 10 seconds, record the volume of gas produced.
5. When the reaction is complete, clean the apparatus as directed by a teacher.
6. Repeat steps 1 to 5 with different concentrations of hydrochloric acid.



## The Effect of Surface Area

For a given mass of a solid, large lumps have smaller surface area to volume ratios than smaller lumps or powders. If a large lump is divided or ground into a powder:

- its total volume stays the same
- the area of exposed surface increases
- the surface area to volume ratio increases



## The Effect of Temperature

The greater the frequency of successful collisions, the greater the rate of reaction, if the temperature of the reaction mixture is increased:

- reactant particles move more quickly
- the energy of the particles increases
- the frequency of successful collisions between reactant particles increases
- the proportion of collisions which are successful increases
- the rate of reaction increases

## The Effect of a Catalyst

- A catalyst is a substance that:
- speeds up the rate of a reaction
  - does not alter the products of the reaction
  - is not chemically changed or used up at the end of the reaction

Only a very small mass of catalyst is needed to increase the rate of a reaction. However, not all reactions are catalysed. Different reactions are catalysed by different chemicals.

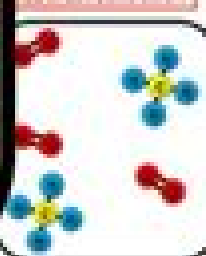


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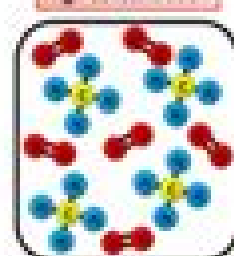
Reversible reaction where the reactants form products that, in turn, form the reactants back. Reversible reactions will reach an equilibrium where the concentrations of the reactants and products will no longer change.

## Concentration

### Low concentration

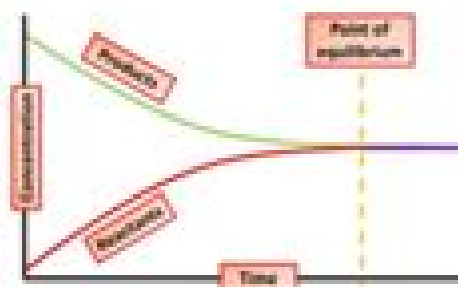


### High concentration



## Equilibrium

Chemical equilibrium is the point at which the concentrations of reactants and products do not change with time. It appears as though the reaction has stopped but in fact the rates of the forward and reverse reactions are equal so reactants and products are being created at the same rate.



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