

CHEMICAL CALCULATIONS

Syllabus reference 8.3.4

For multiple choice questions 1–5 circle the letter corresponding to the most correct answer.

- 1 A mole of H_2 gas contains:
 - a 6.02×10^{23} atoms
 - ☒ b 6.02×10^{23} molecules
 - c 6.02×10^{23} avogadros
 - d 1.2×10^{24} molecules
- 2 The mass of one molecule of nitrogen gas is about:
 - a $\frac{14.0}{2 \times 6.02 \times 10^{23}}$ grams
 - b $2 \times 14.0 \times 6.02 \times 10^{23}$ grams
 - c $\frac{14.0}{6.02 \times 10^{23}}$ grams
 - ☒ d $\frac{2 \times 14.0}{6.02 \times 10^{23}}$ grams
- 3 The mass of 0.100 mole of carbon disulfide (CS_2) is:
 - a 76.0 g
 - ☒ b 7.60 g
 - c 0.76 g
 - d 0.100 g
- 4 A pair of compounds which have the same empirical formula is:
 - a CO_2 ; SO_2
 - b H_2O ; H_2O_2
 - c $\text{C}_2\text{H}_5\text{OH}$; CH_3OH
 - ☒ d $\text{C}_2\text{H}_5\text{OH}$; $(\text{CH}_3)_2\text{O}$
- 5 Hydrogen, oxygen and carbon dioxide gases are contained in three separate flasks. All three flasks have the same volume and all gases have the same temperature and pressure. It follows, from Avogadro's hypothesis and the molar masses of the gases, that:
 - a there are more molecules of carbon dioxide, the heaviest gas of the three
 - b the mass of gas in each flask is the same
 - c there are more molecules of hydrogen, the smallest molecule of the three
 - ☒ d the mass of gas is least for hydrogen, greatest for carbon dioxide.

Conquering Chemistry Prelim Module 2 Ws 5

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