Name: Date:
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#### Student Exploration: Ionic Bonds

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: chemical family, electron affinity, ion, ionic bond, metal, nonmetal, octet rule, shell, valence electron

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

 Nate and Clara are drawing pictures with markers. There are 8 markers in a set. Nate has 9 markers and Clara has 7. What can Nate and Clara do so that each of them has a full set?

If Nate were to give Clara one of his markers, then both would end up with eight.

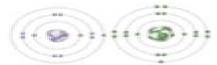
Maggie is sitting at a table with Fred and Florence. Maggie has 10 markers, but Fred and Florence each have only 7 markers. How can they share markers so each has 8?

If Maggie gives Fred and Florence one marker each, they will all have eight markers.

#### Gizmo Warm-up

Just like students sharing markers, atoms sometimes share or swap electrons. By doing this, atoms form bonds. The *lonic Bonds* Gizmo allows you to explore how **ionic bonds** form.

To begin, check that Sodium (Na) and Chlorine (CI) are selected from the menus at right. Click Play ( - to see electrons orbiting the nucleus of each atom. (Note: These atom models are simplified and not meant to be realistic.)



 Each atom consists of a central nucleus and several shells that contain electrons. The outermost electrons are called valence electrons.

How many valence electrons does each atom have? Sodium: 1 Chlorine:

7

Click Pause ( " ). Elements can be classified as metals and nonmetals. Metals do not hold on to their valence electrons very tightly, while nonmetals hold their electrons tightly. Electron affinity is a measure of how tightly the valence electrons are held.

A. Try pulling an electron away from each atom. Based on this experiment, which atom is a metal?

# **Exploration Guide Ionic Bonds Answers**

**Clay Minerals Society** 

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The Origins of Evolutionary Innovations Andreas Wagner, 2011-07-14 The history of life is a nearly four billion year old story of transformative change This change ranges from dramatic macroscopic innovations such as the evolution of wings or eyes to a myriad of molecular changes that form the basis of macroscopic innovations. We are familiar with many examples of innovations qualitatively new phenotypes that provide a critical benefit but have no systematic understanding of the principles that allow organisms to innovate This book proposes several such principles as the basis of a theory of innovation integrating recent knowledge about complex molecular phenotypes with more traditional Darwinian thinking Central to the book are genotype networks vast sets of connected genotypes that exist in metabolism and regulatory circuitry as well as in protein and RNA molecules. The theory can successfully unify innovations that occur at different levels of organization. It captures known features of biological innovation including the fact that many innovations occur multiple times independently and that they combine existing parts of a system to new purposes. It also argues that environmental change is important to create biological systems that are both complex and robust and shows how such robustness can facilitate innovation Beyond that the theory can reconcile neutralism and selectionism as well as explain the role of phenotypic plasticity gene duplication recombination and cryptic variation in innovation Finally its principles can be applied to technological innovation and thus open to human engineering endeavours the powerful principles that have allowed life s spectacular success

Introduction to Modern Inorganic Chemistry Kenneth Malcolm Mackay, Rosemary Ann Mackay, 1989 Iournal of the American Medical Association, 1926 Includes proceedings of the association papers read at the annual sessions and lists of current medical literature **Neither Physics nor Chemistry** Kostas Gavroglu, Ana Simoes, 2011-10-07 The evolution of a discipline at the intersection of physics chemistry and mathematics Quantum chemistry a discipline that is not guite physics not quite chemistry and not quite applied mathematics emerged as a field of study in the 1920s It was referred to by such terms as mathematical chemistry subatomic theoretical chemistry molecular quantum mechanics and chemical physics until the community agreed on the designation of quantum chemistry In Neither Physics Nor Chemistry Kostas Gavroglu and Ana Sim es examine the evolution of quantum chemistry into an autonomous discipline tracing its development from the publication of early papers in the 1920s to the dramatic changes brought about by the use of computers in the 1970s The authors focus on the culture that emerged from the creative synthesis of the various traditions of chemistry physics and mathematics They examine the concepts practices languages and institutions of this new culture as well as the people who established it from such pioneers as Walter Heitler and Fritz London Linus Pauling and Robert Sanderson Mulliken to later figures including Charles Alfred Coulson Raymond Daudel and Per Olov L wdin Throughout the authors emphasize six themes epistemic aspects and the dilemmas caused by multiple approaches social issues including academic politics the impact of textbooks and the forging of alliances the contingencies that arose at every stage of the developments in quantum chemistry

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