

Experiment 34 Prelaboratory Assignment

An Equilibrium Constant

Date _____ Lab Sec. _____ Name _____ Desk No. _____

1. Three parameters affect the absorbance of a sample. Which one is the focus of this experiment?

① Concentration of the solution
② Tube length
③ Wavelength

Concentration is being focused on

2. Experimental Procedure, Part A.1, Table 34.1. A 3.00-mL aliquot of 0.001 M NaSCN is diluted to 25.0 mL with 0.2 M $\text{Fe}(\text{NO}_3)_3$ and 0.1 M HNO_3 .

a. How many moles of SCN^- are present?

$$\text{no. of moles of } \text{SCN}^- = \frac{0.001 \text{ M} \times 3 \text{ mL}}{1000 \text{ mL}} = 3 \times 10^{-6} \text{ mol}$$

- b. If all of the SCN^- is complexed with Fe^{3+} to form FeNCS^{2+} , what is the molar concentration of FeNCS^{2+} ?

$$3 \times 10^{-6} \text{ mol of } \text{SCN}^- \text{ produce } 3 \times 10^{-6} \text{ mol of } \text{FeNCS}^{2+}$$

$$\text{Total volume of diluted solution} = \frac{3 \text{ mL} + 25 \text{ mL}}{1000 \text{ mL}} = 28 \text{ mL}$$

3. Experimental Procedure, Part A.1. For preparing a set of standard solutions of FeNCS^{2+} , the equilibrium molar concentration of FeNCS^{2+} is assumed to equal the initial molar concentration of the SCN^- in the reaction mixture. Why is this assumption valid?

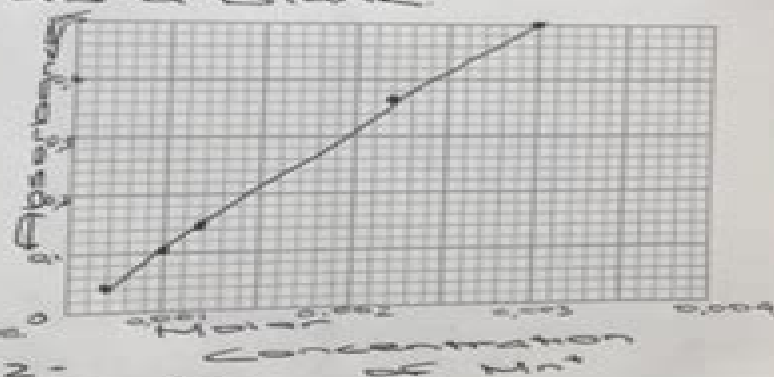
The SCN^- is the limiting reagent and thus when all of SCN^- is added its consumed no more complex formed. Concentration of SCN^- equals to concentration of FeNCS^{2+} .

4. Experimental Procedure, Part A.3. The blank solution used to calibrate the spectrophotometer is 10.0 mL of 0.2 M $\text{Fe}(\text{NO}_3)_3$ diluted to 25.0 mL with 0.1 M HNO_3 . Why is this solution preferred to simply using de-ionized water for the calibration?

The blank prepared above also quashes absorbance value which must be subtracted to the correct value for the complex form when the actual reaction is done. So, plain water cannot be used as a blank.

5. Plot the following data as absorbance versus $[\text{M}^{3+}]$ as a calibration curve:

Absorbance, A	Molar Concentration of M^{3+}
0.045	$3.0 \times 10^{-4} \text{ mol/L}$
0.097	$6.2 \times 10^{-4} \text{ mol/L}$
0.14	$9.0 \times 10^{-4} \text{ mol/L}$
0.35	$2.2 \times 10^{-3} \text{ mol/L}$
0.51	$3.2 \times 10^{-3} \text{ mol/L}$



A test solution showed a percent transmittance (%T) reading of 38.4%T. Interpret the calibration curve to determine the molar concentration of M^{3+} in the test solution.

$$\text{Absorbance} = 2 - 10\log\%T = 2 - 10\log(38.4) = 0.416$$

$$\text{Concentration of unknown solution } (\text{Mn}^{3+}) = 0.416 \times 3.2 \times 10^{-3} / 0.51 = 2.61 \times 10^{-3} \text{ mol/L}$$

Experiment 34 An Equilibrium Constant Answers

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