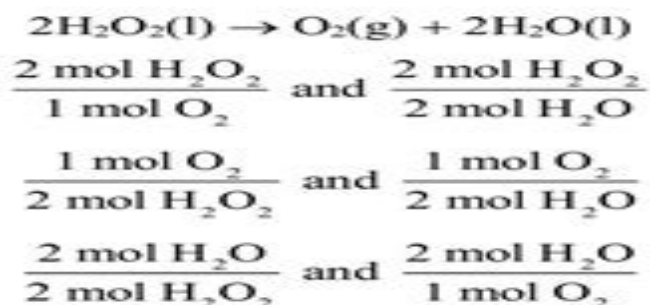


Chapter 11 (continued)



To determine the number of mole ratios that defines a given chemical reaction, multiply the number of species in the equation by the next lower number. Thus, a chemical reaction with three participating species can be defined by six mole ratios ($3 \times 2 = 6$); a chemical reaction with four species can be defined by 12 mole ratios ($4 \times 3 = 12$).

Why learn to write mole ratios? They are the key to calculations that are based on chemical equations. Using a balanced chemical equation, mole ratios derived from the equation, and a given amount of one of the reactants or products, you can calculate the amount of any other participant in the reaction.

Practice Problems

2. Determine all the mole ratios for the following balanced chemical equations.
 - a. $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{NO}(\text{g})$
 - b. $4\text{NH}_3(\text{aq}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{l})$
 - c. $4\text{HCl}(\text{aq}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + 2\text{Cl}_2(\text{g})$

11.2 Stoichiometric Calculations

There are three basic stoichiometric calculations: mole-to-mole conversions, mole-to-mass conversions, and mass-to-mass conversions. All stoichiometric calculations begin with a balanced equation and mole ratios.

► **Stoichiometric mole-to-mole conversion** How can you determine the number of moles of table salt (NaCl) produced from 0.02 moles of chlorine (Cl_2)?

Exercise 11 Stoichiometry 1

**Julie Henderleiter, Charles Trapp, Peter
Atkins, Loretta Jones**

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Congressional Record United States. Congress, 1878

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