

Activity A:

Subatomic particles

Get the Gizmo ready:

- Use the arrows to create an atom with two protons, two neutrons, and two electrons.
- Turn on **Show element name**.



Question: What are the properties of protons, neutrons, and electrons?

1. **Observe:** Turn on **Show element symbol** and **Element notation**. Three numbers surround the element symbol: the **mass number** (A), electrical charge (no number is displayed if the atom is neutral), and the **atomic number** (Z).
2. **Investigate:** Watch how the numbers change as you add or remove particles.



- A. Which number is equal to the number of protons in the atom?
- B. How can you calculate the number of neutrons (N) in an atom?
- C. Which particle (proton, neutron, or electron) has a
Positive charge?
Negative charge?

neutrons

by the number of protons

electron

proton

neutron

3. **Observe:** Turn on **Isotope notation**. Isotope notation shows the number of protons and neutrons in an atom. The mass number (A) is the sum of protons and neutrons. The atomic number (Z) is the number of protons. The electrical charge is shown as a superscript to the right of the element symbol.

- A. What are the mass numbers of carbon?
- B. What are the mass numbers of oxygen?
- C. Carbon has two isotopes: carbon-12 and carbon-13. How many neutrons does each have?

12 neutrons and 13 neutrons

16 neutrons and 17 neutrons

6

4. **Observe:** Turn on **Isotope notation** to answer the following questions.

- A. How many electrons are in a neutral atom of lithium?
- B. How many electrons are in an atom of hydrogen?
- C. What is the mass number of an atom with 3 protons and 3 neutrons?
- D. What is an atom's charge if it has 11 protons and 10 electrons?
- E. How many electrons are in hydrogen?

3

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