

6. **Observe:** Click **Reset**. Move the **Reactant concentration** slider back and forth. What do you notice?

You see more & less reactants & products appear

7. **Predict:** How will increasing the reactant concentration affect the rate of the reaction? Why?

The more reactants, the more products will be created

8. **Gather data:** Make sure the **Temperature** is 200 °C and the **Surface area** is **Maximum**. Use the Gizmo to measure the half-life for each given reactant concentration. (Note that the number of reactant molecules changes with each concentration.) Calculate the means.

Trial	0.4 mol/L	0.8 mol/L	1.2 mol/L	1.6 mol/L	2.0 mol/L
1	3.52	4.16	1.43	1.02	1.08
2	4.03	1.13	1.02	0.56	0.38
Mean half-life	3.77	2.64	1.22	.76	.73

9. **Compare:** If possible, find the mean times for each concentration for your entire class. What is the mean class time for a concentration of 0.4 mol/L? How about for 2.0 mol/L?

Mean for 0.4 mol/L: 3.77

Mean for 2.0 mol/L: .73

10. **Analyze:** What do these results indicate? Still the 2.0 mol/L is the

Explore Learning Gizmo Solubility And Temperature Techer Guide

Thirumalaisamy P.
Velavancorresponding



Explore Learning Gizmo Solubility And Temperature Techer Guide:

Unveiling the Energy of Verbal Artistry: An Psychological Sojourn through **Explore Learning Gizmo Solubility And Temperature Techer Guide**

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