

## Teacher Guide: Boyle's Law and Charles' Law



### Learning Objectives

- observe the effects of pressure, volume, and temperature
- define pressure
- describe how the pressure relates to the volume of a gas at a constant temperature (Boyle's law)
- describe how the temperature relates to the volume of a gas under constant pressure (Charles' law)



### Background

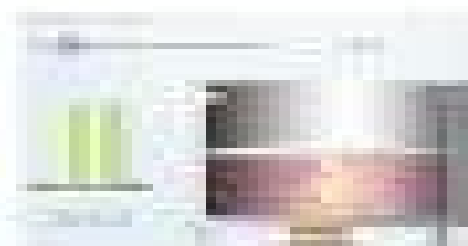
Students often think about "Boyle's" and "Charles" laws as separate concepts.



### Assessments

Boyle's law illustrates how gas molecules move pressure and temperature affect a gas under constant, the volume of a gas changes directly in response to temperature under pressure changes.

In the Boyle's law demonstration, a gas (air) is contained in a cylinder with a piston that can move up and down. The pressure of the gas is constant, and the volume of the gas is constant.



The following questions are designed to assess student understanding:

- Question 1: Students determine the relationship between gas volume and pressure at constant temperature (Boyle's law)
- Question 2: Students determine the relationship between gas volume and temperature at constant pressure (Charles' law)



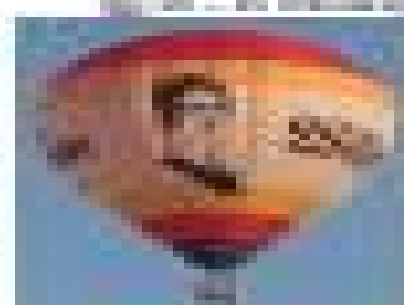
### Suggested Student Response

1. The student activity  
The gas in the cylinder can be heated by increasing the temperature of the cylinder.

When you increase the volume of a gas, the pressure of the gas decreases. This is because the gas molecules are moving faster and are hitting the walls of the cylinder more often. This is why the pressure of the gas decreases when the volume of the gas increases.

When you increase the volume of a gas, the pressure of the gas decreases. This is because the gas molecules are moving faster and are hitting the walls of the cylinder more often. This is why the pressure of the gas decreases when the volume of the gas increases.

Figure 1. Boyle's Law



# Explore Learning Boyles Law Teacher Guide

**N Colangelo**



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