

### School Disco

1. Time?
2. Student Ticket? €?
3. Wear?
4. Address?
5. Food or Drinks?



Saturday Night: 7pm to 11pm

Student: € 8

Must wear a costume

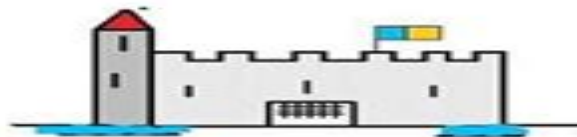
33, Main Street, School Hall  
Cold Drinks available

### Castle Black

1. Open / Sunday?
2. Adult ticket? €?
3. Food?
4. Car Park?
5. Telephone Number?

### Castle Black

Open seven days a week, 9am to 6pm



Adults: € 9 Students: € 5 Children: € 3  
Snacks, cakes, coffee and cold drinks

For more information: Tel: 659244  
Free Car Park

### School Trip

1. Where?
2. Day?
3. Student / Bring?
4. Expensive?
5. Take / Mobile Phone?



Blue Mountains

Tuesday 1<sup>st</sup> April, 9am to 4pm

Students must bring: coat, boots and lunch  
€ 20 per student

No mobile phones

# Cambridge Ket Speaking Cue Cards

**ML Morrison**



## Cambridge Ket Speaking Cue Cards:

*Examining Speaking* Lynda Taylor, 2011-08-25 An up to date review of the relevant literature on assessing speaking

Textile World ,1928      **The Publishers' Trade List Annual** ,1998      Electrical Merchandising Week ,1921      **"The"**  
**Athenaeum** ,1836      **The Quest for Good Governance** Alina Mungiu-Pippidi, 2015-08-25 Why do some societies manage to control corruption so that it manifests itself only occasionally while other societies remain systemically corrupt This book is about how societies reach that point when integrity becomes the norm and corruption the exception in regard to how public affairs are run and public resources are allocated It primarily asks what lessons we have learned from historical and contemporary experiences in developing corruption control which can aid policy makers and civil societies in steering and expediting this process Few states now remain without either an anticorruption agency or an Ombudsman yet no statistical evidence can be found that they actually induce progress Using both historical and contemporary studies and easy to understand statistics Alina Mungiu Pippidi looks at how to diagnose measure and change governance so that those entrusted with power and authority manage to defend public resources      Electrical Merchandising ,1921      **Government Reports Announcements & Index** ,1975-07      Canadian Periodical Index ,1962      The Church Times ,1904      *Author-title Catalog* University of California, Berkeley. Library, 1963      The Kaleidoscope: or, Literary and scientific mirror ,1827

Arts & Humanities Citation Index ,1980 A multidisciplinary index covering the journal literature of the arts and humanities It fully covers 1 144 of the world s leading arts and humanities journals and it indexes individually selected relevant items from over 6 800 major science and social science journals      **Subject Guide to Books in Print** ,1971

*The Billboard* ,1937      Editor & Publisher ,1949 The fourth estate      *Business Week* ,1995      **Canadian Index to Periodicals and Documentary Films** ,1962      **Harvard Business School Bulletin** ,1979      The Saturday Evening Post ,1937 SCC Library has 1974 89 plus scattered issues

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web 1 define each remember oil rig oxidation is loss of e reduction is gain of e a oxidation b reduction c oxidizing agent d reducing agent 2 write half reactions for each of the following atoms or ions label each as oxidation or reduction use the activity series chart a Al b Ba c Br d Ca e Ga f H g H

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web during oxidation and reduction the oxidation numbers of atoms change if an atom is oxidized its oxidation number increases ie it becomes more ve or less ve if an atom is reduced its oxidation number decreases ie it becomes less ve or more ve

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web you should try to answer the questions without referring to your textbook if you get stuck try asking another group for help 1 identify the oxidation state of each element in the following a Na<sub>2</sub>SO<sub>4</sub> b MnO<sub>4</sub><sup>-</sup> c KClO<sub>3</sub> d CH<sub>3</sub>OH 2 identify the oxidation state of nitrogen in the following a NO<sub>3</sub><sup>-</sup> b NO<sub>2</sub> c NO<sub>2</sub>

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web an oxidation reduction redox reaction is one in which one or more atoms change oxidation numbers oxidation occurs when an atom's oxidation state becomes more positive indicating that electrons have been lost reduction occurs when an atom's oxidation state becomes less positive indicating that electrons have been gained

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web describe redox reactions in terms of electron transfer recall that oxidation involves loss of electrons and reduction involves gain of electrons represent simple redox reactions using half equations identify oxidising and reducing agents for simple redox reactions understand the reactivity of metals and metal displacement reaction

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you  $23 \text{ Fe} + \text{Co}^{2+} \rightarrow \text{Fe}^{2+} + \text{Co}$   
oxidation  $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$  reduction  $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$   $24 \text{ 3 Ag} + \text{Ni}^{2+} \rightarrow 3 \text{ Ag}^{+} + \text{Ni}$  oxidation reduction  $25 \text{ Cu} + \text{Pb}^{2+} \rightarrow \text{Cu}^{2+} + \text{Pb}$  oxidation reduction  $26 \text{ O}_2$

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web in each of the following redox reactions indicate the elements that undergo changes in oxidation number and the magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent  $\text{a } \text{I}_2 + \text{O}_5 \rightarrow \text{I} + \text{O}_2$   
 $5\text{Co} + \text{I}_2 + 5\text{Co}^{2+} + \text{I}^{-} + 2\text{O}_5 \rightarrow 5\text{Co} + \text{I}_2 + 5\text{Co}^{2+} + \text{I}^{-} + 2\text{O}_2$

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also  $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$   $\text{Sr} \rightarrow \text{Sr}^{2+}$  oxidized reducing agent  $\text{O}_2 \rightarrow \text{O}^{2-}$  reduced ox ag  $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$   $\text{Li} \rightarrow \text{Li}^{+}$  oxidized red ag  $\text{S} \rightarrow \text{S}^{2-}$  reduced ox ag  $3\text{C}_2\text{S} + \text{Br}_2$