

# Chemical Equations and Reactions

Quiz  
CH 8  
Section #1

## Section Quiz: Describing Chemical Reactions

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- B 1. Which of the following is true regarding the law of conservation of mass?
- It dictates that the number of molecules on each side of a chemical equation must be the same.
  - It dictates that the number of atoms of each element must be the same on both sides of a chemical equation.
  - It states that the mass of the reactants must remain constant in order for a chemical reaction to proceed.
  - It does not apply to chemical reactions.
- D 2. Which of the following may indicate that a chemical reaction has occurred?
- release of energy as light
  - a color change
  - gas bubble formation
  - All of the above
- D 3. Which of the following is a product in the reaction described by the word equation below?
- $$\text{iron} + \text{copper(II) sulfate} \rightarrow \text{iron(II) sulfate} + \text{copper}$$
- iron(II) sulfate
  - copper
  - iron
  - Both (a) and (b)
- C 4. What does  $\text{CH}_4$  represent in the balanced chemical equation shown below?
- $$\text{C} + 2\text{H}_2 \rightarrow \text{CH}_4$$
- one atom of  $\text{CH}_4$
  - five atoms of  $\text{CH}_4$
  - one molecule of  $\text{CH}_4$
  - 1 g of  $\text{CH}_4$
- C 5. Which of the following is not true regarding a properly written chemical equation?
- It must be balanced.
  - It must contain proper chemical formulas.
  - It must show the formation of a precipitate or a gas.
  - It should represent known data.

# Ch 8 Modern Chemistry Study Guide

**Charlotte W. Pratt, Kathleen Cornely**



## Ch 8 Modern Chemistry Study Guide:

**Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03  
The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

**Fundamentals of Chemistry, Study Guide** James E. Brady, John R. Holum, 1988-04-20  
This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up-to-date treatments of important and current topics. Emphasizes descriptive chemistry, chemical reactions, and properties while maintaining a solid treatment of chemical principles. Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises. Incorporates many pedagogical aids; each chapter begins with a brief table of contents, and each section begins with a preview of topics covered. Chapters include frequent margin comments, figures, and photographs.

**A Guide to Modern Chemistry** John F. Baxter, Jay A. Young, 1959

**A Guide to Undergraduate Science Course and Laboratory Improvements** National Science Foundation (U.S.). Directorate for Science Education, 1979

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