Guizizz

Ch. 18 (Chemical Equilibrium) 20 Questions NAME:

CLASS:

DATE:

CO(g) + H₂O(g)

CO₂(g) + H₂(g)
 This is the expression for the equilibrium constant for the reaction shown above.

$$K_{eq} = \frac{[CO_2]^2[H_2]}{[CO][H_2O]^2}$$

$$K_{eq} = \frac{[CO_2][H_2]}{[CO][H_2O]}$$

$$K_{eq} = \frac{[CO][H_20]}{[CO_2][H_2]}$$

2. The hydrogen sulfide produced as a byproduct of petroleum refinement can be used to produce elemental sulfur: 2H₂S(g) + SO₂(g) ≒ 3S(l) + 2H₂O(g) What is the equilibrium constant expression for this reaction?

$$K_{eq} = \frac{[H_2O]^2}{[H_2S]^2[SO_2]}$$

$$K_{eq} = \frac{[S]^{3}[H_{2}O]^{2}}{[H_{2}S]^{2}[SO_{2}]}$$

$$K_{eq} = \frac{[H_2S]^2[SO_2]}{[H_2S]^2}$$

$$K_{eq} = \frac{[H_2O]}{[H_2S][SO_2]}$$

- 3. What can be determined if the ion product (Q_{sp}) is compared to the solubility product constant (K_{sp})?
- A the law of mass action

B whether a substance will precipitate

equilibrium

- the volume of the solution
- 4. The value of any equilibrium constant is correct only at
- A a specific volume.

B a specific pressure.

a specific concentration.

- a specific temperature.
- 5. A system reaches chemical equilibrium when

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SB Merriam



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