MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

Crystalline solids ______

A) exist only at very low temperatures

B) have their particles arranged randomly

C) have highly ordered structures

D) are usually very soft

E) exist only at high temperatures

In liquids, the attractive intermolecular forces are ______

 A) strong enough to hold molecules relatively close together but <u>not</u> strong enough to keep molecules from moving past each other

B) not strong enough to keep molecules from moving past each other

C) very weak compared with kinetic energies of the molecules

D) strong enough to hold molecules relatively close together

E) strong enough to keep the molecules confined to vibrating about their fixed lattice points

3) Based on molecular mass and dipole moment of the five compounds in the table below, which should have the highest boiling point?

Substance		Dipole Moment (D)
Propane, CH3CH2CH3	44	0.1
Dimethylether, CH3OCH3	46	1.3
Methylchloride, CH3Cl	50	1.9
Acetaldehyde, CH3CHO	44	2.7
Acetonitrile, CH3CN	41	3.9

A) CH3OCH3

B) CH₃CN

C) CH₃CHO

D) CH3CL

E) CH3CH2CH3

4) The intermolecular force(s) responsible for the fact that CH4 has the lowest boiling point in the set

CH4, SiH4, GeH4, SnH4 is/are _____

A) mainly hydrogen bonding but also dipole-dipole interactions

B) hydrogen bonding

C) mainly London-dispersion forces but also dipole-dipole interactions

D) dipole-dipole interactions

E) London dispersion forces

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