

Answers to Chapter 9 Study Questions

Study Questions:

1. Why is it important that we eat food and breathe oxygen? Food is broken down through digestion to small organic molecules such as glucose. Glucose and oxygen are reactants in cellular respiration, a process by which cells generate energy (ATP and heat) and produce the waste products (CO₂ and H₂O). Without this energy, cells cease to function and organisms die.
2. Why do plants need CO₂ and H₂O as well as light energy? These molecules are reactants for photosynthesis – the process where plants generate organic molecules (glucose) and oxygen. Plants used these molecules to make ATP via cellular respiration.
3. What types of molecules are used within the body to generate energy? Monosaccharides (glucose), fatty acids, glycerol, and amino acids. These molecules are oxidized to generate CO₂ and water. In general, why does glucose have more potential energy than CO₂? Glucose has more chemical bonds that can be rearranged through chemical reactions. The rearrangement of some of these bonds releases energy – thus the bonds in glucose contain much more potential energy. CO₂ does not have chemical bonds that can be easily rearranged and thus it has very little potential energy.
4. Is the oxidation of glucose an endergonic or an exergonic reaction? Exergonic. Why? The change in free energy is -686 kcal/mol. How is the energy that is generated through the oxidation of glucose use in the cell? About 34% of the energy is used to generate ATP. The rest is released as heat. Heat can be use for thermoregulation or dissipated from the body through sweat.
5. Chemical reactions involve the transfer of electrons between molecules. What happens when electrons are relocated between molecules? Energy is released when electrons are transferred. This is why molecules such as methane and glucose have so much potential energy.
6. What are oxidation-reduction reactions? Reactions that involve the partial or complete transfer of one or more electron from one reactant to another. Use the reaction below to answer the following questions



What is oxidation and which one of the reactants in the above equation is the reducing agent (electron donor)? Oxidation is the pull of shared electrons away from one atom. In the equation, electrons in methane are equally shared between C and H because these atoms have similar electronegativities. When methane combines with 2 O₂ to produce CO₂ and H₂O, the electrons are pulled away from C and towards O when CO₂ is formed because O is more electronegative. Thus the carbon in CH₄ loses shared electrons and it is the reducing agent (electron donor).

What is reduction and which one of the reactants in the above equation is the oxidizing agent (electron acceptor)? Reduction is the pull of shared electrons towards another atom. When methane reacts with 2 O₂ molecules to form CO₂ and H₂O, the electrons that were shared equally in O₂ are now pulled towards O in H₂O. Thus the oxygen in H₂O “gains” the shared electrons and it is the oxidizing agent (electron acceptor).

Why is oxygen such a powerful oxidizing agent? It is very electronegative and will readily accept electrons. What type of redox reaction releases the most chemical energy? Reactions where electrons are relocated to oxygen.

Ch 9 Study Guide Answers

Annelies Wilder-Smith

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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Table of Contents Ch 9 Study Guide Answers

1. Understanding the eBook Ch 9 Study Guide Answers
 - The Rise of Digital Reading Ch 9 Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ch 9 Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ch 9 Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ch 9 Study Guide Answers
 - Personalized Recommendations
 - Ch 9 Study Guide Answers User Reviews and Ratings
 - Ch 9 Study Guide Answers and Bestseller Lists
5. Accessing Ch 9 Study Guide Answers Free and Paid eBooks
 - Ch 9 Study Guide Answers Public Domain eBooks
 - Ch 9 Study Guide Answers eBook Subscription Services
 - Ch 9 Study Guide Answers Budget-Friendly Options
6. Navigating Ch 9 Study Guide Answers eBook Formats

- ePub, PDF, MOBI, and More
- Ch 9 Study Guide Answers Compatibility with Devices
- Ch 9 Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ch 9 Study Guide Answers
 - Highlighting and Note-Taking Ch 9 Study Guide Answers
 - Interactive Elements Ch 9 Study Guide Answers
- 8. Staying Engaged with Ch 9 Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ch 9 Study Guide Answers
- 9. Balancing eBooks and Physical Books Ch 9 Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ch 9 Study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ch 9 Study Guide Answers
 - Setting Reading Goals Ch 9 Study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ch 9 Study Guide Answers
 - Fact-Checking eBook Content of Ch 9 Study Guide Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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