

Name: [KEY] Period: _____ Date: _____

I. Dilution Problems $M_1V_1 = M_2V_2$

1. A student wishes to prepare 250. mL of a 0.15 M NaOH solution from a 6.00M NaOH stock solution. What volume of the stock solution should she use?

$$M_2 = 0.15 \text{ M} \quad M_1 = 6.00 \text{ M} \\ V_2 = 250. \text{ mL} \quad V_1 = ? \\ (6.00 \text{ M})(x \text{ mL}) = (0.15 \text{ M})(250. \text{ mL}) \\ x = \boxed{6.25 \text{ mL}} \text{ stock solution}$$

2. How much water must be added to a 2.50 M HCl solution to obtain 500. mL of a 0.100M HCl solution?

$$M_1 = 2.50 \text{ M} \quad M_2 = 0.100 \text{ M} \\ V_1 = ? \quad V_2 = 500. \text{ mL} \\ (2.50 \text{ M})(x \text{ mL}) = (0.100 \text{ M})(500. \text{ mL}) \\ x = 20 \text{ mL} \\ 500 \text{ mL} - 20 \text{ mL} = \boxed{480 \text{ mL}} \\ \text{(total volume)} \quad \text{(stock solution volume)} \quad \text{volume of water added}$$

3. What is the final volume of a 0.15 M solution prepared from 25.0 mL of a 6.0 M sodium acetate solution?

$$M_1 = 6.0 \text{ M} \quad M_2 = 0.15 \text{ M} \\ V_1 = 25.0 \text{ mL} \quad V_2 = ? \\ (6.0 \text{ M})(25.0 \text{ mL}) = (0.15 \text{ M})(x \text{ mL}) \\ x = \boxed{1,000 \text{ mL}}$$

II. Molarity Problems $\text{Molarity (M)} = \frac{\text{mol}}{\text{L}}$

4. What is the molarity of a solution that contains 12.5g $\text{NaC}_2\text{H}_3\text{O}_2$ (Sodium Acetate) dissolved in 0.250 L?

$$12.5 \text{ g NaC}_2\text{H}_3\text{O}_2 \times \frac{1 \text{ mol of NaC}_2\text{H}_3\text{O}_2}{82.04 \text{ g NaC}_2\text{H}_3\text{O}_2} = 0.1524 \text{ moles of NaC}_2\text{H}_3\text{O}_2 \\ M = \frac{(0.1524 \text{ mol})}{(0.250 \text{ L})} = \boxed{0.609 \text{ M}}$$

5. What mass of citric acid (molar mass = 192.14 g/mol) would be needed to make 0.10 L of 5.00 M citric acid solution?

$$5.00 \text{ M} = \frac{x \text{ moles}}{0.10 \text{ L}} \\ x = 0.50 \text{ moles} \\ 0.50 \text{ moles citric acid} \times \frac{192.14 \text{ g citric acid}}{1 \text{ mol citric acid}} = 96.07 \text{ g rounded} \\ \boxed{96.1 \text{ g of citric acid}}$$

6. What mass of sodium hydroxide is needed to make 0.500 L of 2.50 M NaOH?

$$2.50 \text{ M} = \frac{x \text{ moles}}{0.500 \text{ L}} \\ x = 1.25 \text{ moles} \\ 1.25 \text{ moles NaOH} \times \frac{40 \text{ g NaOH}}{1 \text{ mol NaOH}} = \boxed{50 \text{ g of NaOH}}$$

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Alexander Cockburn



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