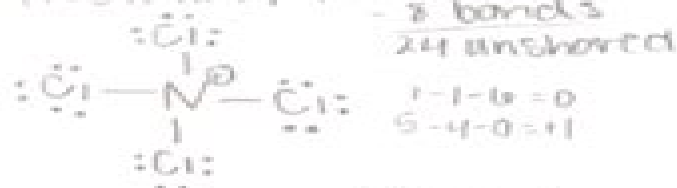


# Chapter 9 Practice Problems

1) Draw Lewis Structures for each.

$$\text{NCl}_3: (1 \times 5) + (4 \times 7) - 1 = 32 \text{ electrons}$$



PBr<sub>3</sub>

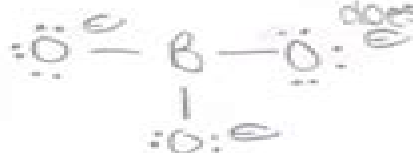
$$1 \times 5 + 3 \times 7 = 26 \text{ electrons}$$

5 - 3 - 2 = 0  
7 - 1 - 6 = 0

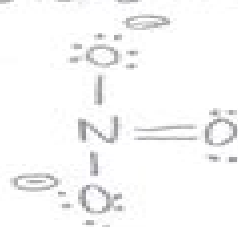


BO<sub>3</sub><sup>3-</sup>

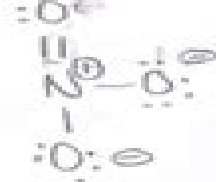
Boron is an exception  
does not make octet



$$\text{NO}^+ : 1 \times 5 + 5 \times 6 + 1 = 32$$



$$\begin{array}{r} 32 \\ - 24 \\ \hline 8/4 \text{ bonds} \end{array}$$

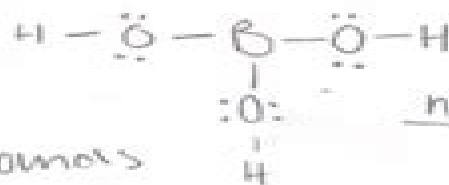
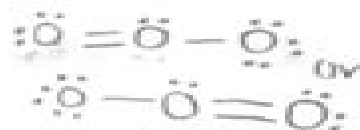


$$\text{SF}_2: 1 \times 6 + 1 \times 7 = 20 \text{ electrons}$$

2 bonds  
18 unshared



$$\text{O}_3: 1 \times 6 + 1 \times 6 + 1 \times 6 = 18 \text{ electrons}$$



no charges

resonance  
arrow

1A	2A		3A	4A	5A	6A	7A	8A
H	He		B	C	N	O	F	Ne
Li	Be		Al	Si	P	S	Cl	Ar
Na	Mg		Ga	Ge	As	Se	Br	Kr
K	Ca	Sc	In	Sn	Sb	Te	I	Xe
Rb	Sr	Y	Tl	Pb	Bi	Po	At	Rn
Cs	Ba	La	Pb	Bi	Po	At	Rn	

# Ch 9 Practice Problems Chemistry Answers

**Greg Curran**



## Ch 9 Practice Problems Chemistry Answers:

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

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Purvis-Roberts, 2022-10-21 SOLVE Problems in Environmental Science delivers up a robust set of engaging quantitative problems geared toward students in guided problem solving groups and Environmental Science courses In response to repeated requests for more problems in environmental science Katie Purvis Roberts Claremont McKenna Pitzer and Scripps Colleges and Tom Spiro University of Washington authors of Chemistry of the Environment with a team of experienced environmental science teachers have developed SOLVE Problems in Environmental Science This sleek and affordable stand alone problems book serves up a broad array of quantitative problems addressing real world issues in an approachable fashion Requiring only algebra and a basic understanding of general chemistry SOLVE is designed for use in traditional Environmental Science courses as well as in student centered guided problem solving courses Worked problems are followed by practice problems with brief answers that allow students to check their work With this text your students will use their reasoning ability to tackle and solve problems ranging from global warming to GMOs An Instructor s Manual with detailed solutions is also available to adopting professors **CliffsNotes Chemistry Practice Pack** Charles Henrickson, 2010-02-08

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