

Name: KEY

Section: \_\_\_\_\_

**Chapter 7 Practice Worksheet:  
Covalent Bonds and Molecular Structure**

1) How are ionic bonds and covalent bonds different?

Ionic bonds result from the transfer of electrons from one atom to another; Covalent bonds result from two atoms sharing electrons.

2) Describe the relationship between the length of a bond and the strength of that bond.

Strength of a bond increases as the bond gets shorter (inverse relationship)

3) Identify the type(s) of bond(s) found in the following molecules:

- a.  $\text{CCl}_4$       covalent
- b.  $\text{Li}_2\text{O}$       ionic
- c.  $\text{NF}_3$       covalent
- d.  $\text{CaSO}_4$       ionic and covalent
- e.  $\text{SO}_2$       covalent
- f.  $\text{Mg}(\text{OH})_2$       ionic and covalent

4) Define electronegativity.

Electronegativity is the ability for an atom in a molecule to attract electrons to itself.

5) Use electronegativity values to place the following elements in **increasing** order: F, N, H, P, Si, C, O

$\text{Si} < \text{C} \approx \text{P} \approx \text{H} < \text{N} < \text{O} < \text{F}$

6) Determine if the bond between atoms in each example below is nonpolar covalent, polar covalent, or ionic.

- |                  |              |                 |              |
|------------------|--------------|-----------------|--------------|
| a. $\text{H}_2$  | <u>npc</u>   | e. $\text{NF}$  | <u>pc</u>    |
| b. $\text{PCl}$  | <u>pc</u>    | f. $\text{MgO}$ | <u>ionic</u> |
| c. $\text{F}_2$  | <u>npc</u>   | g. $\text{CH}$  | <u>npc</u>   |
| d. $\text{NaBr}$ | <u>ionic</u> | h. $\text{HCl}$ | <u>pc</u>    |

# Ch 7 Wksts Cp Guerinchemistry

**RJ Shavelson**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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## **Table of Contents Ch 7 Wksts Cp Guerinchemistry**

1. Understanding the eBook Ch 7 Wksts Cp Guerinchemistry
  - The Rise of Digital Reading Ch 7 Wksts Cp Guerinchemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying Ch 7 Wksts Cp Guerinchemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ch 7 Wksts Cp Guerinchemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ch 7 Wksts Cp Guerinchemistry
  - Personalized Recommendations

- Ch 7 Wksts Cp Guerinchemistry User Reviews and Ratings
- Ch 7 Wksts Cp Guerinchemistry and Bestseller Lists
- 5. Accessing Ch 7 Wksts Cp Guerinchemistry Free and Paid eBooks
  - Ch 7 Wksts Cp Guerinchemistry Public Domain eBooks
  - Ch 7 Wksts Cp Guerinchemistry eBook Subscription Services
  - Ch 7 Wksts Cp Guerinchemistry Budget-Friendly Options
- 6. Navigating Ch 7 Wksts Cp Guerinchemistry eBook Formats
  - ePub, PDF, MOBI, and More
  - Ch 7 Wksts Cp Guerinchemistry Compatibility with Devices
  - Ch 7 Wksts Cp Guerinchemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ch 7 Wksts Cp Guerinchemistry
  - Highlighting and Note-Taking Ch 7 Wksts Cp Guerinchemistry
  - Interactive Elements Ch 7 Wksts Cp Guerinchemistry
- 8. Staying Engaged with Ch 7 Wksts Cp Guerinchemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ch 7 Wksts Cp Guerinchemistry
- 9. Balancing eBooks and Physical Books Ch 7 Wksts Cp Guerinchemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ch 7 Wksts Cp Guerinchemistry
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ch 7 Wksts Cp Guerinchemistry
  - Setting Reading Goals Ch 7 Wksts Cp Guerinchemistry
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ch 7 Wksts Cp Guerinchemistry
  - Fact-Checking eBook Content of Ch 7 Wksts Cp Guerinchemistry

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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