

Chapter 10: (see also Problem Setways Page)

Date _____

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Computation involving mole fractions and the given information involving the formulae at bottom for the unknowns and plugging in the information/numbers and solving and writing the answer legibly and

1. A sample of 0.800 moles of gas is collected in a container at volume of 0.010 L. What is the pressure of the gas at 200 K if the gas is at 20 °C? (R=8.314 J/mol·K)

$$P = \frac{nRT}{V} = \frac{(0.800 \text{ mol})(8.314 \text{ J/mol}\cdot\text{K})(200 \text{ K})}{(0.010 \text{ L})} = 1.33 \times 10^5 \text{ Pa}$$

2. A 0.010 L sample of gas is collected at 200 K. The pressure of the gas is 1.33 x 10⁵ Pa. What is the volume of the gas at 20 °C?

$$V = \frac{nRT}{P} = \frac{(0.800 \text{ mol})(8.314 \text{ J/mol}\cdot\text{K})(200 \text{ K})}{(1.33 \times 10^5 \text{ Pa})} = 0.010 \text{ L}$$

3. A sample of 0.800 moles of gas is collected in a container at volume of 0.010 L. What is the pressure of the gas at 200 K if the gas is at 20 °C? (R=8.314 J/mol·K)

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4. A 0.010 L sample of gas is collected at 200 K. The pressure of the gas is 1.33 x 10⁵ Pa. What is the volume of the gas at 20 °C?

$$V = \frac{nRT}{P} = \frac{(0.800 \text{ mol})(8.314 \text{ J/mol}\cdot\text{K})(200 \text{ K})}{(1.33 \times 10^5 \text{ Pa})} = 0.010 \text{ L}$$

5. A sample of 0.800 moles of gas is collected in a container at volume of 0.010 L. What is the pressure of the gas at 200 K if the gas is at 20 °C? (R=8.314 J/mol·K)

$$P = \frac{nRT}{V} = \frac{(0.800 \text{ mol})(8.314 \text{ J/mol}\cdot\text{K})(200 \text{ K})}{(0.010 \text{ L})} = 1.33 \times 10^5 \text{ Pa}$$

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$$V = \frac{nRT}{P} = \frac{(0.800 \text{ mol})(8.314 \text{ J/mol}\cdot\text{K})(200 \text{ K})}{(1.33 \times 10^5 \text{ Pa})} = 0.010 \text{ L}$$

Ch 11 Gases Review Answers

PT Brinkman

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the PT Brinkman text.

Ch 11 Gases Review Answers:

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- Chapter 3: Ch 11 Gases Review Answers in Everyday Life
- Chapter 4: Ch 11 Gases Review Answers in Specific Contexts
- Chapter 5: Conclusion

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