

**PROCEDURE****CHECK ALL GLASSWARE!!!**

For this experiment, use the table below (Table 1) to help you prepare your mixtures.

**Table 1.** Reagents for determining the rate orders of  $\text{MnO}_4^-$ ,  $\text{H}_2\text{C}_2\text{O}_4$ , and  $\text{H}^+$ .

Reagents	Mixture 1	Mixture 2	Mixture 3	Mixture 4
0.50 M $\text{H}_2\text{C}_2\text{O}_4$	1.00 mL	1.00 mL	2.00 mL	1.00 mL
0.10 M $\text{H}_2\text{SO}_4$	1.00 mL	1.00 mL	1.00 mL	2.00 mL
Distilled $\text{H}_2\text{O}$	2.00 mL	3.00 mL	1.00 mL	1.00 mL
<hr/>				
0.0050 M $\text{KMnO}_4$	2.00 mL	1.00 mL	2.00 mL	2.00 mL
Total	6.00 mL	6.00 mL	6.00 mL	6.00 mL

**NOTE:** Contamination can be a real problem for this experiment. Pipettes meant for a particular solution should be used only for that solution. The test tube should be clean and dry before kinetics run.

**BEFORE STARTING ON PART A:** Start your hot water bath for part B. Fill the beaker full of water.

**Part A: Determining the Rate Law**

# Clarkson Gpc Chemistry 1212 Lab Manual

**LL Leslie**



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