

Exercise 1: The First Law of Thermodynamics

1. A system is defined as a collection of particles. The system is isolated if no energy or matter can enter or leave the system. The system is closed if energy can enter or leave the system, but matter cannot. The system is open if both energy and matter can enter or leave the system. The system is a control volume if it is a region in space through which matter flows, but energy does not flow across the boundary of the control volume.

2. The first law of thermodynamics states that the change in the internal energy of a system is equal to the heat added to the system minus the work done by the system.

$$\Delta U = Q - W$$

where ΔU is the change in internal energy, Q is the heat added to the system, and W is the work done by the system.

3. The internal energy of a system is a state function, which means that its value depends only on the state of the system, not on the path taken to reach that state. The internal energy of a system can be measured by calorimetry, which is the study of heat and work. The internal energy of a system can also be calculated from the equation of state, which relates the pressure, volume, and temperature of a system.

4. The first law of thermodynamics can be applied to a variety of systems, including gases, liquids, and solids. It can be used to calculate the change in internal energy of a system, the heat added to a system, and the work done by a system.

5. The first law of thermodynamics is a statement of the conservation of energy. It states that energy cannot be created or destroyed, only transformed from one form to another.

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