

A concentrated phosphoric acid solution is 85.5% H_3PO_4 by mass and has a density of 1.69 g/mL at 25°C.

- What is the molarity of H_3PO_4 ?
- What is the mole fraction of urea, $\text{CO}(\text{NH}_2)_2$, in a solution prepared by dissolving 8.2 g of urea in 45.1 g of methanol, CH_3OH ?
- How will an understanding of concentration help you in your healthcare career?

1. What is the molarity of H_3PO_4 ? Understanding Molarity, it is a way to express the concentration of a solution, specifically how many moles of a solute are present in one liter of solution.

Given:

Mass percent of H_3PO_4 = 85.5%

Density of the solution = 1.69 g/mL

STEPS:

a. Calculate the mass of H_3PO_4 in 1 liter (1000 mL) of solution.

Mass of solution = Density $\times$ Volume

$$1.69 \text{ g/mL} \times 1000 \text{ mL} = 1690 \text{ g}$$

b. Calculate the mass of H_3PO_4 in this solution.

$$\text{Mass of } \text{H}_3\text{PO}_4 = 0.855 \times 1690 = 1444.95 \text{ g}$$

c. Convert mass of H_3PO_4 to moles.

$$\text{The molar mass of } \text{H}_3\text{PO}_4 \text{ is } 3(1.01) + 30.97 + 4(16.00) = 97.99 \text{ g/mol}$$

$$\text{d. Calculate molarity: Moles of } \text{H}_3\text{PO}_4 = 1444.95 \text{ g} / 97.99 \text{ g/mol} = 14.74 \text{ M}$$

2. What is the mole fraction of urea, $\text{CO}(\text{NH}_2)_2$, in a solution prepared by dissolving 8.2 g of urea in 45.1 g of methanol, CH_3OH ?

Given: Mass of urea = 8.2 g and Mass of methanol = 45.1 g

a. Convert mass of urea to moles

$$\text{The molar mass of urea } (\text{CO}(\text{NH}_2)_2) \text{ is approx. } 12.01 + (2 \times 14.01) + (4 \times 1.01) = 60.06 \text{ g/mol}$$

$$\text{Moles of urea} = 8.2 \text{ g} / 60.06 \text{ g/mol} = 0.136 \text{ mol}$$

b. Convert mass of methanol to moles.

$$\text{The molar mass of methanol } (\text{CH}_3\text{OH}) \text{ is approx. } 12.01 + (4 \times 1.01) + 16.00 = 32.05 \text{ g/mol}$$

$$\text{Moles of methanol} = 45.1 \text{ g} / 32.05 \text{ g/mol} = 1.40 \text{ mol}$$

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R Barnett



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