Post-lab questions:

- Using the volumes of HSO₃ solution added for each of the three titrations you conducted (A, B, and C), calculate the number of moles HSO₃ required to reach the endpoint in each titration.
- Using the moles of HSO₃⁻ calculated in #1 above, calculate the number of moles of MnO₄⁻ titrated according to the balanced equations you submitted in the prelab exercises.
- Using the moles of MnO₄ you calculated in #2 above, calculate the starting concentration of permanganate ion for each of the three titrations.
- 4. Write the electron configurations for Mn, Mn2+, Mn4+, Mn6+, and Mn7+.
- Sodium bismuthate (NaBiO₃) oxidizes MnCl₂ to MnO₄ and is consequently reduced to Bi(OH)₃. How many grams of MnCl₂ will be oxidized by 7.5 g of NaBiO₃?
- 6. Balance the following oxidation-reduction reactions:
 - a. Mn²⁺ + BiO₃⁻ → MnO₄⁻ + Bi³⁺ (in acid)
 - b. NO₃ + I₂ → NO₂ + IO₃ (in acid)

- c. ClO₃⁻ + S²- → Cl⁻ + S (in base)
- d. Mo(OH)₃ + HO₂⁻ → MoO₄² (in base)
- 7. In each of the titrations performed in this experiment, the end point was apparent due to a color change in the material being oxidized or reduced. Suppose that none of the materials involved in the redox reaction are colored. Suggest a way one could determine the endpoint in such a case. (HINT: Think about acid-base titrations.)

Marco Cascella

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