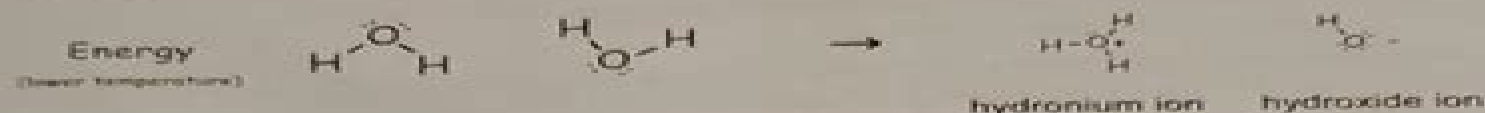


Conceptual Chemistry

Chapter 10: Acids and Bases

Supercritical Water

In an endothermic process, water reacts with water to form hydronium and hydroxide ions:



With a higher temperature, there is more thermal energy available and this favors more product:



1. What then happens to the pH? (circle one)

a) It decreases

b) It increases

c) It remains the same

Please explain:

2. With higher temperature, the concentration of hydronium ions increases. What happens to the concentration of hydroxide ions? Please explain:

3. This solution is (circle one)

a) acidic

b) neutral

c) basic

Please explain:

At temperatures above 100°C water exists as a gas (water vapor). Apply pressure and you can literally squeeze the gas back into the liquid phase. However, above a temperature of 374°C and a pressure of 218 atmospheres, the densities of the liquid and gaseous phases of water become equal, which allows the two phases to blend into a single phase called a "supercritical fluid". Amazingly, in its supercritical phase, water takes on a new set of properties. Notably, many nonpolar materials dissolve quite well in supercritical water. This includes oxygen, O₂, and organic chemicals, such as those found in toxic wastes, which can be combusted right within the supercritical water.

4. What are the primary chemicals produced from the combustion of organic compounds within a supercritical water combustion chamber?

5. Would supercritical water combustion be useful for destroying radioactive wastes?

6. Why is supercritical water so potentially damaging to the interiors of a supercritical water combustion chamber?

Conceptual Chemistry Practice Page Answers

**Donald J. Wink, Sharon Fetzer-
Gislason, Sheila McNicholas**



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