

Concentration and Molarity II: Dilution and Evaporation PhET-Chemistry Lab



Introduction: In the first part of this lab, you learned about the actions of solid salts added to water to form solutions of various concentrations. In this section, you will use a powerful and simple formula to determine the concentration of a solution when it has been diluted with additional water or made more concentrated due to evaporation of water.

Some handy equations for you to define:

Molarity _____

Dilution _____

Evaporation _____

Procedure: PHET-Poly with the Chem. Ph Chemistry Ph Concentration 100%

$$M_1V_1 = M_2V_2$$

Part I: Dilution

At this point you should be familiar with the Concentration simulation. Because some time may have passed, take a few minutes to re-learn the simulation. Pay particular attention to the effect of evaporation and addition of water to the solution's concentration.



How does the concentration change as additional water is added? _____

Why? _____



How does the concentration change as evaporation occurs? _____

Why is this? _____

Does evaporation change the concentration of a saturated solution? _____

Why is this? _____



Using the concentrated solution spigot, add a 10 Liter of Drink Mix to an empty beaker. What is the concentration?

_____ Is this solution saturated? _____ How do you know? _____

Fill the beaker with another 10 Liter of water. What is the new concentration? _____

Complete the table below, using 100% Cobalt chloride in an empty beaker, writing the concentration in the boxes provided.

Dilute .25L of spigot solution	.25L spigot+.25L water	.25L spigot + .50 L water	.25L spigot + .75 L water

Repeat the exercises, using 100% Potassium dichromate in an empty beaker.

Dilute .25L of spigot solution	.25L spigot+.25L water	.25L spigot + .50 L water	.25L spigot + .75 L water

What do you notice about the concentration change as each addition of .25L of water is added to the concentrated spigot solution? _____

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Answers Key

LL Leslie

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape with a gradient, partially overlapping the bar's end.

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Table of Contents Concentration And Molarity Phet Chemistry Labs Answers Key

1. Understanding the eBook Concentration And Molarity Phet Chemistry Labs Answers Key
 - The Rise of Digital Reading Concentration And Molarity Phet Chemistry Labs Answers Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Concentration And Molarity Phet Chemistry Labs Answers Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Concentration And Molarity Phet Chemistry Labs Answers Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Concentration And Molarity Phet Chemistry Labs Answers Key
 - Personalized Recommendations
 - Concentration And Molarity Phet Chemistry Labs Answers Key User Reviews and Ratings
 - Concentration And Molarity Phet Chemistry Labs Answers Key and Bestseller Lists
5. Accessing Concentration And Molarity Phet Chemistry Labs Answers Key Free and Paid eBooks
 - Concentration And Molarity Phet Chemistry Labs Answers Key Public Domain eBooks
 - Concentration And Molarity Phet Chemistry Labs Answers Key eBook Subscription Services
 - Concentration And Molarity Phet Chemistry Labs Answers Key Budget-Friendly Options

6. Navigating Concentration And Molarity Phet Chemistry Labs Answers Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Concentration And Molarity Phet Chemistry Labs Answers Key Compatibility with Devices
 - Concentration And Molarity Phet Chemistry Labs Answers Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Concentration And Molarity Phet Chemistry Labs Answers Key
 - Highlighting and Note-Taking Concentration And Molarity Phet Chemistry Labs Answers Key
 - Interactive Elements Concentration And Molarity Phet Chemistry Labs Answers Key
8. Staying Engaged with Concentration And Molarity Phet Chemistry Labs Answers Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Concentration And Molarity Phet Chemistry Labs Answers Key
9. Balancing eBooks and Physical Books Concentration And Molarity Phet Chemistry Labs Answers Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Concentration And Molarity Phet Chemistry Labs Answers Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Concentration And Molarity Phet Chemistry Labs Answers Key
 - Setting Reading Goals Concentration And Molarity Phet Chemistry Labs Answers Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Concentration And Molarity Phet Chemistry Labs Answers Key
 - Fact-Checking eBook Content of Concentration And Molarity Phet Chemistry Labs Answers Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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