

## CHM 3342: Test 1 Review

Chapters covered: 1, 2, 3, and 4

Test date: 2/24/2023 (Friday)

*You must use my lecture slides to guide your reading of the textbook. The lecture slides may contain some extra stuff not present in the textbook. Pay close attention to the highlighted points (bold or colored texts) in my PowerPoint slides. I will not use any of the "end-of-the-chapter" questions except those I have selected and mentioned on the chapter's last slide.*

*There are two types of questions in the exam. The first is conceptual ones. The second is applying the concepts to solve problems related to real structures. The latter type of questions will have heavier weight in the exam.*

*Please bring a blue scantron for the exam.*

### Chapter 1. Structure and Bonding

- 1.1 to 1.3: Atomic Structure. The shape of orbitals is **not required** except for  $s$  and  $p$  orbitals. Know how to write the ground-state electron configuration of common elements.
- 1.4: Know how to draw the Lewis and Kekule structures of a given molecule.
- 1.5-1.10: Know the difference between  $\sigma$  and  $\pi$  bonds. Understand the  $sp^3$ ,  $sp^2$ , and  $sp$  hybridizations and their implication on the shape of a molecule. You should be able to determine the hybridization of C, N, O, S, and halogens in an organic molecule.
- 1.11: **Not required.**
- 1.12: Know how to deduce the full structure and formula from a skeletal structure.

### Chapter 2. Polar Covalent Bonds; Acids and Bases

- 2.1-2.2: Know the general order of electronegativity in the Periodic Table. The specific value for each element is **not required**. Know how to determine one molecule would have dipole moments or not. Know the unit and formula for dipole moment, however the calculation is **not required**.
- 2.3: Know to calculate formal charges (examples in lecture slides).
- 2.4-2.6: Know how to draw the resonance forms of a given structure. Be able to use curved arrow to show the relevance of different resonance forms.
- 2.7-2.10: Know the definitions of  $K_a$ ,  $pK_a$  and their correlation with acid and base strength. Know how to draw the conjugate base of an acid, and the conjugate acid of an base. **Calculation of  $K_a$  and  $pK_a$  is not required.**
- 2.11: Know the definition and show which component is Lewis acid and which one is Lewis base in a given acid-base reaction.
- 2.12: Know the features of noncovalent interactions: dipole-dipole forces, dispersion forces, and hydrogen bonds.

# Chm 02 Study Guide And Notebook

**S Ben Porath**



## **Chm 02 Study Guide And Notebook:**

## Decoding **Chm 02 Study Guide And Notebook**: Revealing the Captivating Potential of Verbal Expression

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