

12

STOICHIOMETRY

Practice Problems

In your notebook, solve the following problems.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

1. An apple pie needs 10 large apples, 2 cups (top and bottom), and 1 tablespoon of cinnamon. Write a balanced equation that fits this situation. How many apples are needed to make 20 pies?
2. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate, $\text{KClO}_3(\text{s})$. Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
3. Using the equation from problem 2, how many moles of oxygen are produced from 14 moles of potassium chlorate?
4. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from 3.8×10^{23} molecules of oxygen? How many moles of water are produced from 22.5 moles of oxygen?

SECTION 12.2 CHEMICAL CALCULATIONS

1. Calculate the number of moles of hydrogen chloride produced from 10 moles of hydrogen.



2. Calculate the number of moles of chlorine needed to form 14 moles of iron(II) chloride.



3. Calculate the number of grams of nitrogen dioxide that are produced from 4 moles of nitric oxide.



4. Calculate the mass of oxygen produced from the decomposition of 75.0 g of potassium chlorate.



5. Calculate the mass of silver needed to react with chlorine to produce 94 g of silver chloride. Write a balanced equation first.

6. How many liters of carbon monoxide at STP are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



7. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of dinitrogen monoxide. Assume all gases are at the same conditions of temperature and pressure.



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Sebastian Brünink



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1. Understanding the eBook Core Teaching Resources Chemistry Answer Sheet
 - The Rise of Digital Reading Core Teaching Resources Chemistry Answer Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Core Teaching Resources Chemistry Answer Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Core Teaching Resources Chemistry Answer Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Core Teaching Resources Chemistry Answer Sheet
 - Personalized Recommendations
 - Core Teaching Resources Chemistry Answer Sheet User Reviews and Ratings
 - Core Teaching Resources Chemistry Answer Sheet and Bestseller Lists
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 - Core Teaching Resources Chemistry Answer Sheet Enhanced eBook Features
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- Adjustable Fonts and Text Sizes of Core Teaching Resources Chemistry Answer Sheet
 - Highlighting and Note-Taking Core Teaching Resources Chemistry Answer Sheet
 - Interactive Elements Core Teaching Resources Chemistry Answer Sheet
8. Staying Engaged with Core Teaching Resources Chemistry Answer Sheet
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Core Teaching Resources Chemistry Answer Sheet
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 - Creating a Diverse Reading Collection Core Teaching Resources Chemistry Answer Sheet
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- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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- Setting Reading Goals Core Teaching Resources Chemistry Answer Sheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Core Teaching Resources Chemistry Answer Sheet
- Fact-Checking eBook Content of Core Teaching Resources Chemistry Answer Sheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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